

Name: _____ Date: _____

Student Exploration: Nuclear Decay

Vocabulary: alpha particle, atomic number, beta particle, daughter product, gamma ray, isotope, mass number, nuclear decay, positron, radioactive, subatomic particle

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

The chart below gives the locations, charges, and approximate masses of three **subatomic particles**. The approximate mass of each particle is given in universal mass units (u).

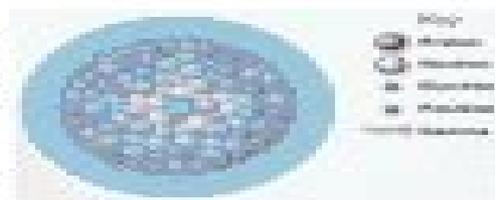
Particle	Location	Charge	Approximate mass
Proton	Nucleus	1^+	1 u
Neutron	Nucleus	0	1 u
Electron	Orbitals	1^-	0 u

- The **mass number** of an atom is equal to the sum of protons and neutrons in the nucleus. A helium atom has 2 protons and 2 neutrons. What is the mass number of this atom? _____
- The **atomic number** of an element is equal to the number of protons in each atom of the element. All helium atoms have 2 protons. What is the atomic number of helium? _____

Gizmo Warm-up

While most atoms are stable, some are **radioactive**, which means that they have a tendency to undergo spontaneous **nuclear decay**. The decay of radioactive atoms generally results in the emission of particles and/or energy.

Several types of nuclear decay can be explored with the *Nuclear Decay Gizmo™*. On the Gizmo, check that **Alpha decay** and **Uranium** are selected.



- Click **Play** (▶) and then click **Pause** (⏸) when the **alpha particle** is clearly visible. What is an alpha particle made of? _____
- Click **Play** and observe. Besides the alpha particle, what else is emitted from the nucleus during alpha decay? _____

Gamma rays are energetic electromagnetic waves; they are often emitted in nuclear decay.

Explore Learning Nuclear Decay Gizmo Answer Key

**Donald Douglas Burgess, McMaster
University. Department of Chemistry**

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Burgess, McMaster University. Department of Chemistry, 1974 *The Determination of Nuclear Decay Schemes by Absorption and Coincidence Measurements* John Patrick Lonergan, 1951

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