

# Planning an Experiment into Solubility Answers

## 1. Background Information:

When a substance is dissolved into a liquid, we call the substance dissolving a **solute**, the liquid a **solvent**, and the mixture formed is called a **solution**.

We are going to plan an experiment to see how much solute dissolves at different temperatures, this is known as **solubility**. If we keep adding the solute and no more dissolves, then the solution is fully **saturated**.

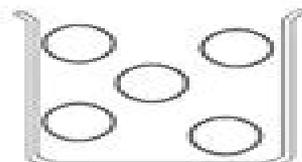
## 2. Method:

- Clear a suitable working area, tie long hair up and put on safety goggles.
- Fill a glass beaker with 25ml of sugar (solvent) at room temperature.
- Add 10g of sugar (solute) at a time, noting down how much you have added.
- Stir gently with a glass rod, applying the same amount of force each time.
- Keep adding the solute until no more dissolves and the solution is fully saturated and record the total amount of grams of solute added.
- Repeat steps b to d with hotter water. The temperatures we will use are 50°C and 100°C. Students completing the higher ability worksheet may have used different temperatures.

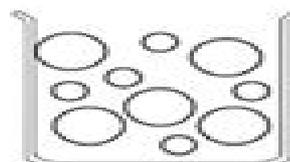
## 3. How the Particles Look Inside

Draw two beakers showing the particles inside of them.

Beaker A (water only)



Beaker B (water and solute)



## 4. Risk Assessment

Complete the table to consider any safety issues in this practical. The first one has been given as an example.

Hazard	Harm	How You Will Prevent Injuries
glass beaker	Could break and cut skin.	Keep floor clear and hold beaker securely; place in middle of table; wear goggles.
hot water	Could scold or burn skin.	Keep floor clear and hold beaker securely; place in middle of table; cover arms if possible.
water	Could spill on floor and slip over.	Mop up any spillages straight away and inform others.
thermometer	Could break and cut skin.	Hold thermometer securely and return when not in use so it cannot roll off the table; use a glass stirring rod to mix the sugar.



# Explore Learning Solubility And Temperature Answers

**Annelies Wilder-Smith**



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