

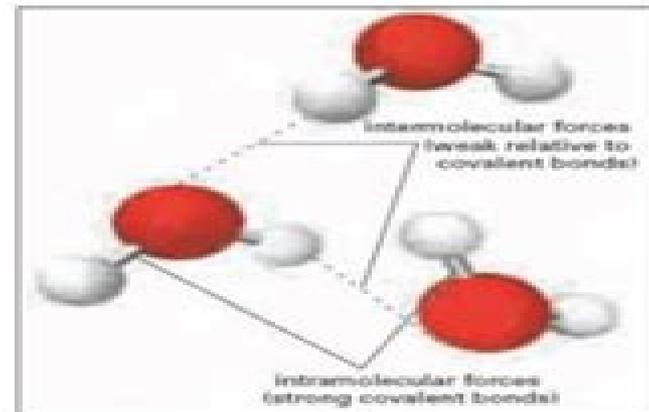
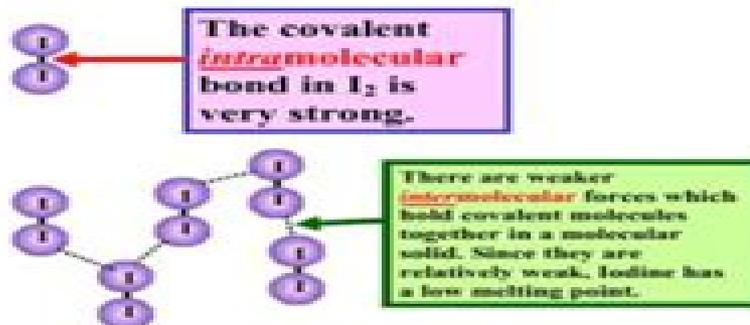
# Intermolecular Forces

## Intermolecular Forces:

- Forces between molecules
- Responsible for the state of matter: solid, liquid, or gas
- affect the melting and boiling points of compounds as well as the solubility of one substance in another
- weaker than covalent bonds within molecules, since molecular compounds melt easily (melting doesn't break the bond between atoms, but between the molecules themselves!)
  - o example: melting ice to water to steam will break apart the intermolecular forces and change state, but not decompose into  $H_2$  and  $O_2$

## Intramolecular Force (Force-Force attraction):

- Forces within molecules
- the attractive force between the valence electrons of one atom and the nucleus of another atom would be considered intermolecular forces
- Covalent bonds are examples of intramolecular forces



**Figure 3.20** Strong intramolecular forces (covalent bonds) hold the atoms in molecules together. Relatively weak intermolecular forces act between molecules.

# Exploring Intermolecular Forces Lab Answers

**M Mark**



## **Exploring Intermolecular Forces Lab Answers:**

## **Exploring Intermolecular Forces Lab Answers** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Exploring Intermolecular Forces Lab Answers**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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