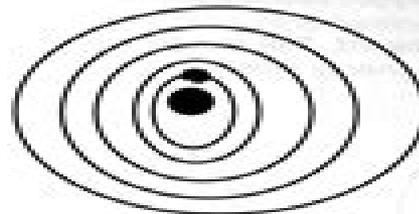


In this activity, we will use a representation of the atom in which a central nucleus containing the protons and neutrons is surrounded by circles that represent the energy levels electrons can occupy.

- 1) Draw an atom including a nucleus and five energy levels that electrons could occupy. Use a dot to represent an electron at the lowest energy level.



One way an atom emits light (photons) occurs when an electron drops down from a high energy level (also referred to as an excited state) to a lower energy level (the lowest energy level is referred to as the ground state.)

- 2) Will an atom emit light if all of the atom's electrons are in the ground state? Explain your reasoning.

An atom cannot emit light if all the atom's electrons are in the ground state because there would be no higher energy level to lower energy levels to drop down to.

- 3) In which case does an atom emit more energy (circle one)?

Case A:

An electron drops down from the first excited state to the ground state.

Case B:

An electron drops down from the third excited state to the ground state.

Explain your reasoning.

An electron dropping down from the third excited state to the ground state would emit more energy because from the third excited state to the ground state, there is more difference in energy which is needed for an atom to emit energy.

Light And Atoms Lecture Tutorial For Astronomy S

JG Myers



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