

# Ionic Compounds and Formula Worksheet

Write the formula for each of the following compounds.

|                       |                                    |                         |                                    |
|-----------------------|------------------------------------|-------------------------|------------------------------------|
| 1. Magnesium oxide    | <b>MgO</b>                         | 2. Sodium fluoride      | <b>NaF</b>                         |
| 3. Aluminum nitride   | <b>Al<sub>2</sub>N<sub>3</sub></b> | 4. Potassium sulfide    | <b>K<sub>2</sub>S</b>              |
| 5. Lithium iodide     | <b>LiI</b>                         | 6. Calcium bromide      | <b>CaBr<sub>2</sub></b>            |
| 7. Beryllium oxide    | <b>BeO</b>                         | 8. Nickel chloride      | <b>NiCl</b>                        |
| 9. Magnesium nitride  | <b>Mg<sub>3</sub>N<sub>2</sub></b> | 10. Aluminum sulfide    | <b>Al<sub>2</sub>S<sub>3</sub></b> |
| 11. Calcium phosphide | <b>Ca<sub>3</sub>P<sub>2</sub></b> | 12. Zinc fluoride       | <b>ZnF</b>                         |
| 13. Sodium oxide      | <b>Na<sub>2</sub>O</b>             | 14. Lithium chloride    | <b>LiCl</b>                        |
| 15. Zinc oxide        | <b>ZnO</b>                         | 16. Silver chloride     | <b>AgCl</b>                        |
| 17. Barium bromide    | <b>BaBr<sub>2</sub></b>            | 18. Potassium phosphide | <b>K<sub>3</sub>P</b>              |
| 19. Silver oxide      | <b>Ag<sub>2</sub>O</b>             |                         |                                    |

Write the names for the following compounds.

|                                    |                            |                                    |                            |
|------------------------------------|----------------------------|------------------------------------|----------------------------|
| 1. Li <sub>2</sub> O               | <b>Lithium oxide</b>       | 2. CaF <sub>2</sub>                | <b>Calcium fluoride</b>    |
| 3. AlCl <sub>3</sub>               | <b>Aluminum chloride</b>   | 4. K <sub>3</sub> P                | <b>Potassium phosphide</b> |
| 5. MgS                             | <b>Magnesium sulfide</b>   | 6. ZnF <sub>2</sub>                | <b>Zinc fluoride</b>       |
| 7. CaO                             | <b>Calcium oxide</b>       | 8. AgCl                            | <b>Silver chloride</b>     |
| 9. KBr                             | <b>Potassium bromide</b>   | 10. Al <sub>2</sub> S <sub>3</sub> | <b>Aluminum sulfide</b>    |
| 11. BeF                            | <b>Beryllium fluoride</b>  | 12. BaBr <sub>2</sub>              | <b>Barium bromide</b>      |
| 13. Na <sub>3</sub> N              | <b>Sodium nitride</b>      | 14. CaS                            | <b>Calcium sulfide</b>     |
| 15. Al <sub>2</sub> O <sub>3</sub> | <b>Aluminum oxide</b>      | 16. KI                             | <b>Potassium iodide</b>    |
| 17. Na <sub>2</sub> S              | <b>Sodium sulfide</b>      | 18. ZnO                            | <b>Zinc oxide</b>          |
| 19. Mg <sub>3</sub> P <sub>2</sub> | <b>Magnesium phosphide</b> |                                    |                            |

# Formulae Of Ionic Compounds Answer Key

**CH Cherryholmes**



**Formulae Of Ionic Compounds Answer Key:**

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