

(b) Here, $\lambda = 198 \text{ nm} = 198 \times 10^{-9} \text{ m}$

$$\omega_0 = 2.5 \text{ eV}$$

Energy of the incident photon

$$\begin{aligned} E &= \frac{hc}{\lambda} = \frac{6.63 \times 10^{-34} \times 3 \times 10^8}{198 \times 10^{-9}} \text{ J} \\ &= \frac{6.63 \times 3}{198} \times \frac{10^{-17}}{1.6 \times 10^{-19}} \text{ eV} \\ &= 6.28 \text{ eV} \end{aligned}$$

$$\begin{aligned} \therefore \text{Energy of the ejected electrons} &= 6.28 - 2.50 = 3.78 \text{ eV} \\ &= 3.78 \times 1.6 \times 10^{-19} \text{ J} \end{aligned}$$

If V is the stopping potential of the electrons,

$$\begin{aligned} \text{eV} &= 3.78 \times 1.6 \times 10^{-19} \text{ J} \\ 1.6 \times 10^{-19} \times V &= 3.78 \times 1.6 \times 10^{-19} \end{aligned}$$

$$\therefore V = 3.78 \text{ Volts}$$

(c) Here, $E_n = -\frac{13.6}{n^2} \text{ eV}$

$$\therefore E_2 = -\frac{13.6}{4} \text{ eV} = -3.4 \text{ eV}$$

and $E_1 = -\frac{13.6}{1} = -13.6 \text{ eV}$

Energy released when the transition takes place from $n = 2$ to $n = 1$

$$E = E_2 - E_1 = -3.4 - (-13.6) = 10.2 \text{ eV} = 10.2 \times 1.6 \times 10^{-19} \text{ J}$$

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