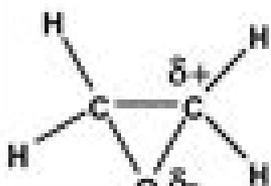


Question	Answer	Marks	Guidance
2 (a) (i)	$2C_2H_6 + O_2 \rightarrow 2C_2H_4O$ ✓	1	ALLOW molecular formulae OR correct structural OR displayed OR skeletal formula OR mixture of the above (as long as unambiguous) ALLOW correct multiples, including fractions, of this equation IGNORE state symbols DO NOT ALLOW [O]
	(ii) $C_2H_6 + 3O_2 \rightarrow 2CO_2 + 2H_2O$ ✓	1	ALLOW molecular formulae OR correct structural OR displayed OR skeletal formula OR mixture of the above (as long as unambiguous) ALLOW correct multiples of this equation IGNORE state symbols DO NOT ALLOW [O]
(b) (i)		1	Only one carbon atom needs to have the correct partial charge DO NOT ALLOW partial charges on hydrogen atoms
	(ii) Movement of an electron pair ✓	1	ALLOW movement of a lone pair OR movement of a bond ALLOW movement of two electrons
	(iii) Heterolytic ✓ Both electrons (in the bond) go to the same atom OR (bond breaks) to make a cation and (a lone pair on the oxygen atom) OR bond pair becomes a lone pair on oxygen ✓	2	MARK INDEPENDENTLY ALLOW one atom gets none of the bonded electrons DO NOT ALLOW both electrons go to a molecule DO NOT ALLOW makes a positive and a negative ion because in this example this is <i>not</i> true
	(iv) It donates a pair of electrons ✓	1	ALLOW donates a lone pair DO NOT ALLOW it donates electrons
	(v) idea that H^+ ion is used in step 1 AND made in step 4 ✓	1	ALLOW H^+ ion is used at the start AND made at the end IGNORE overall H^+ is not used up in the mechanism

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DJ Losen



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