

- Here are three isotopes of an element:  $^{12}_6\text{C}$        $^{13}_6\text{C}$        $^{14}_6\text{C}$ 
  - The element is: **Carbon**
  - The number 6 refers to the: **Atomic Number**
  - The numbers 12, 13, and 14 refer to the **different mass numbers**
  - How many protons and neutrons are in the first isotope? **P: 6    N: 6**
  - How many protons and neutrons are in the second isotope? **P: 6    N: 7**
  - How many protons and neutrons are in the third isotope? **P: 6    N: 8**

2. Complete the following chart:

Isotope name	atomic #	mass #	# of protons	# of neutrons	# of electrons
Potassium-37	19	37	19	18	19
Oxygen-17	8	17	8	9	8
uranium-235	92	235	92	143	92
uranium-238	92	238	92	146	92
boron-10	5	10	5	5	5
boron-11	5	11	5	6	5

3. How do you calculate the mass number of an element?

**Mass # = protons + neutrons**

- If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
  - Silicon-14
  - Silicon-28
  - Silicon-31**
  - Silicon-17
- If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
  - Iodine-53
  - Iodine-127
  - Iodine-54
  - Iodine-107**

# Isotopes Practice Set Answer Key

**Edith Hinkley Quimby**



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