

Mayya's Equilibrium Study Guide



K_c is the Equilibrium Constant

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$



K_p is the Equilibrium Constant
when pressures are given

$$K_p = \frac{P_c^c P_d^d}{P_a^a P_b^b}$$

$$K_p = K_c (RT)^{\Delta n}$$

Where $R=0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$, T = temperature in K,
 Δn = moles of gas products - moles of gas reactants

How To Predict The Direction Of Reaction

$$Q = \frac{[C]^c \text{noneq} [D]^d \text{noneq}}{[A]^a \text{noneq} [B]^b \text{noneq}}$$

If $Q > K$, the reaction will go to the left
If $Q < K$, the reaction will go to the right
If $Q = K$, the reaction is at equilibrium

ICE CHART

I = Initial (initial concentrations given in the problem)
C = Change (in terms of X, including coefficients)
E = Equilibrium (I+C)

Make an ICE Chart when you are asked to find equilibrium concentrations or pressures

Equation Stoichiometry & Equilibrium Constant

- Reverse an equation = Take reciprocal of K ($1/K$)
- Multiply an equation by a constant = Raise K to the power of that constant
- Add equations = Multiply their K 's

Equilibrium And Ksp Study Guide

Silberberg



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

[Chemical Principles Study Guide/Solutions Manual](#) John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

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