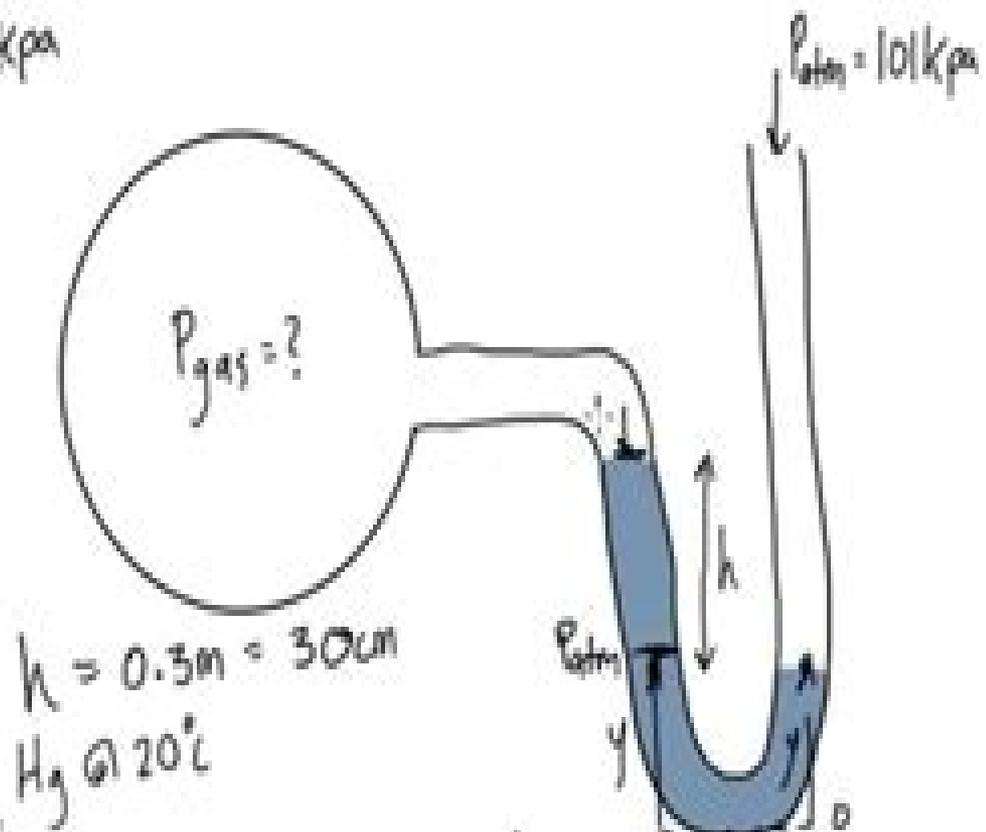


$$\begin{aligned}
 P_{\text{gas}} &= P_{\text{atm}} + \rho g h \\
 &= 101 \text{ kPa} + (13,550 \frac{\text{kg}}{\text{m}^3}) (9.81 \text{ m/s}^2) (0.3 \text{ m}) \\
 &= 101 \text{ kPa} + 39.877 \text{ kPa}
 \end{aligned}$$

$$P_{\text{gas}} = 141 \text{ kPa}$$



$h = 0.3 \text{ m} = 30 \text{ cm}$
 Hg @ 20°C

$$\begin{aligned}
 P_{\text{gas}} &= P_{\text{atm}} - \rho g h \\
 &= 101 \text{ kPa} - (13,550) (9.81) (0.3) \\
 &= 101 - 39.877
 \end{aligned}$$

$$P_{\text{gas}} = 61 \text{ kPa}$$

Manometer Example Problems

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