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Friday 27 May 2016 – Morning

AS/GCE CHEMISTRY A

FP21/01 – Atoms, Bonds and Groups

Candidates answer on the Question Paper

OCR supplied materials:

- Data Sheet for Chemistry A inserted

Other materials required:

- Scientific calculator

Duration: 1 hour

Candidate surname		Candidate surname	
Centre number		Candidate number	

INSTRUCTIONS TO CANDIDATES

- This event will be found inside this document.
- Write your name, centre number and candidate number in the boxes above. Please write clearly and in capital letters.
- Use black ink. Ball-point may be used for graphs and diagrams only.
- Answer all the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- Write your answer to each question in the space provided. If additional answer space is required, you should use the (extra) page(s) at the end of this document. The question number(s) must be clearly shown.
- Do not write in the bar codes.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- Where you see this icon you will be awarded marks for the quality of written communication in your answer. This means for example you should:
 - ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear;
 - organise information clearly and coherently, using specialist vocabulary when appropriate;
 - use only a scientific calculator.
- A copy of the Data Sheet for Chemistry A is provided as an insert with this question paper.
- You are allowed to show all the steps in any calculations.
- The total number of marks for this paper is 56.
- This document consists of 16 pages. Any blank pages are indicated.

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2

6d) Calculate the amount, in mol, of nitrogen atoms in 5.0×10^3 nitrogen molecules.

Give your answer in standard form.

amount of nitrogen atoms = _____ mol [1]

6e) N_2O_5 is an unstable oxide of nitrogen that decomposes in a redox reaction.

$N_2O_5(g) \rightarrow NO(g) + NO_2(g)$

6i) State the oxidation number of nitrogen in each oxide in the table below.

Oxide	Oxidation number of nitrogen
N_2O_5	
NO	
NO_2	

[1]

6ii) Name this type of redox reaction.

In your answer you should use appropriate technical terms spelled correctly.

[1]

6f) N_2O_5 reacts with water to form an acid as the only product. This reaction is **not** a redox reaction. The empirical formula of the acid formed is the same as the molecular formula.

6i) State what is meant by the term molecular formula.

2

Answer all the questions.

1 Nitrogen is the most common gas in the atmosphere.

6a) Atoms of nitrogen consist of protons, neutrons and electrons.

Complete the table below.

Particle	Relative mass	Relative charge	Position within the atom
Proton			
Neutron			
Electron			shell

[4]

6b) The electrons in the second shell of a nitrogen atom are found in an s-orbital and three p-orbitals.

6i) State, in words, the 3D shape of an s-orbital and a p-orbital.

s-orbital _____

p-orbital _____

[2]

6ii) Describe the relative energies of the 2s-orbital and each of the three 2p-orbitals in a nitrogen atom.

[4]

6c) Draw a dot-and-cross diagram to show the bonding in a nitrogen molecule.

Show outer electrons only.

[5]

2 This question is about halogens.

6a) Solid chlorine and solid bromine have a similar structure.

Name this structure.

[1]

6b) The intermolecular attractions in halogens are van der Waals' forces.

6i) Explain how van der Waals' forces arise between halogen molecules.

[4]

6ii) The boiling points of chlorine and bromine are shown in the table.

Halogen	Boiling point / °C
chlorine	-34
bromine	59

Explain why bromine has a higher boiling point than chlorine.

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Table of Contents June 2014 F321 01 Question Paper

1. Understanding the eBook June 2014 F321 01 Question Paper
 - The Rise of Digital Reading June 2014 F321 01 Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2014 F321 01 Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2014 F321 01 Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2014 F321 01 Question Paper
 - Personalized Recommendations
 - June 2014 F321 01 Question Paper User Reviews and Ratings
 - June 2014 F321 01 Question Paper and Bestseller Lists
5. Accessing June 2014 F321 01 Question Paper Free and Paid eBooks

- June 2014 F321 01 Question Paper Public Domain eBooks
 - June 2014 F321 01 Question Paper eBook Subscription Services
 - June 2014 F321 01 Question Paper Budget-Friendly Options
6. Navigating June 2014 F321 01 Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2014 F321 01 Question Paper Compatibility with Devices
 - June 2014 F321 01 Question Paper Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2014 F321 01 Question Paper
 - Highlighting and Note-Taking June 2014 F321 01 Question Paper
 - Interactive Elements June 2014 F321 01 Question Paper
 8. Staying Engaged with June 2014 F321 01 Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2014 F321 01 Question Paper
 9. Balancing eBooks and Physical Books June 2014 F321 01 Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2014 F321 01 Question Paper
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine June 2014 F321 01 Question Paper
 - Setting Reading Goals June 2014 F321 01 Question Paper
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of June 2014 F321 01 Question Paper
 - Fact-Checking eBook Content of June 2014 F321 01 Question Paper
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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