

Student Exploration: Covalent Bonds

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|---|--|---|
| <p>Activity B: Building molecules</p> | <p><u>Get the Gizmo ready:</u></p> <ul style="list-style-type: none"> • Click Reset. • Turn off Show Lewis diagram. • Select Oxygen. |  |
|---|--|---|

Question: How do atoms share more than one pair of electrons?

1. **Observe:** Like fluorine and most other elements, oxygen atoms are most stable with a full complement of eight valence electrons.

A. How many valence electrons does each oxygen atom have now? 6

B. How many more electrons does each oxygen atom need to be stable? 2

2. **Form a bond:** Drag electrons back and forth until the molecule of oxygen (O_2) is stable. Click **Check** to confirm your molecule is stable. Take a snapshot and paste the image into your document (don't forget to label it " O_2 ").

How many *pairs* of shared electrons are there in a stable molecule of oxygen? 1

3. **Draw a diagram:** Draw a Lewis diagram of the oxygen molecule in the space below at left. To check your work, turn on **Show Lewis diagram**. Draw the correct diagram on the right.

Practice diagram: $\cdot\ddot{O} = \ddot{O}\cdot$

Actual: $\cdot\ddot{O} = \ddot{O}\cdot$

4. **Practice:** Create covalent bonds and stable molecules for the remaining substances. Take a snapshot of each completed molecule and add it to your document. Draw Lewis diagrams for each one. (As above, draw the diagram on your own before checking your work.)

Nitrogen



Ammonia



Chlorine



Methane



Water



Formaldehyde



(Activity B continued on next page)

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Padhraic Smyth



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Table of Contents Explore Learning Covalent Bonds Answer Key

1. Understanding the eBook Explore Learning Covalent Bonds Answer Key
 - The Rise of Digital Reading Explore Learning Covalent Bonds Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Covalent Bonds Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Covalent Bonds Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Covalent Bonds Answer Key

- Personalized Recommendations
 - Explore Learning Covalent Bonds Answer Key User Reviews and Ratings
 - Explore Learning Covalent Bonds Answer Key and Bestseller Lists
5. Accessing Explore Learning Covalent Bonds Answer Key Free and Paid eBooks
 - Explore Learning Covalent Bonds Answer Key Public Domain eBooks
 - Explore Learning Covalent Bonds Answer Key eBook Subscription Services
 - Explore Learning Covalent Bonds Answer Key Budget-Friendly Options
 6. Navigating Explore Learning Covalent Bonds Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Covalent Bonds Answer Key Compatibility with Devices
 - Explore Learning Covalent Bonds Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Covalent Bonds Answer Key
 - Highlighting and Note-Taking Explore Learning Covalent Bonds Answer Key
 - Interactive Elements Explore Learning Covalent Bonds Answer Key
 8. Staying Engaged with Explore Learning Covalent Bonds Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Covalent Bonds Answer Key
 9. Balancing eBooks and Physical Books Explore Learning Covalent Bonds Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Covalent Bonds Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Explore Learning Covalent Bonds Answer Key
 - Setting Reading Goals Explore Learning Covalent Bonds Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Explore Learning Covalent Bonds Answer Key

- Fact-Checking eBook Content of Explore Learning Covalent Bonds Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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