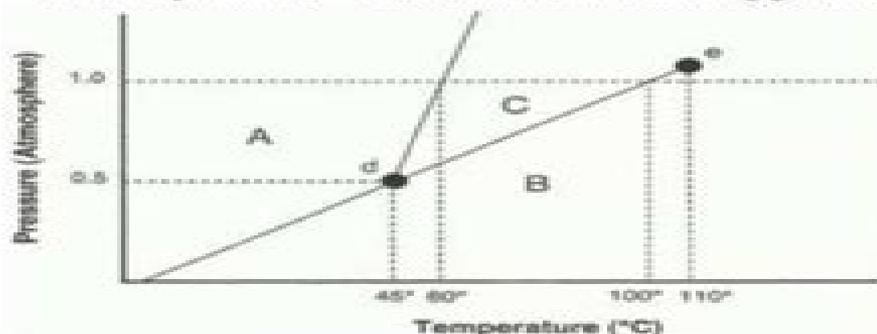


PHASE DIAGRAM REVIEW WORKSHEET

Part A – Generic Phase Diagram.

Answer questions 1 – 10 in relation to the following generic phase diagram.



1. Which section represents the solid phase? A
2. What section represents the liquid phase? C
3. What section represents the gas phase? B
4. What letter represents the triple point? d

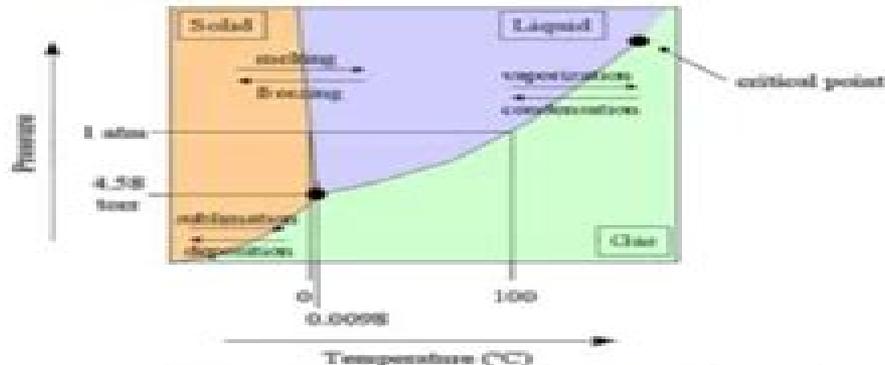
In your own words, what is the definition of a triple point?

answers will vary – s, l, g exist at equilibrium

5. What is this substance's normal melting point, at 1 atmosphere of pressure? 60°C
6. What is this substance's normal boiling point, at 1 atmosphere of pressure? 100°C
7. Above what temperature is it impossible to liquefy this substance, no matter what the pressure? 110°C
8. At what temperature and pressure do all three phases coexist? 0.5 atm, 45°C
9. At a constant temperature, what would you do to cause this substance to change from the liquid phase to the solid phase? increase the pressure
10. What does sublimation mean? to go from a solid to a gas directly

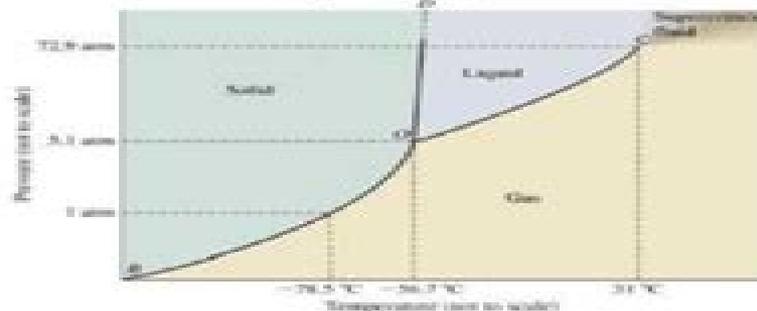
Part B – Phase Diagram for Water.

11. At a pressure of 1 atmosphere, what is the normal freezing point of water? 0°C
12. What is the normal boiling point of water, at one atmosphere of water? 100°C
13. People in Albuquerque live approximately 5,500 feet above sea level, which means the normal atmospheric pressure is less than 1 atm. In Albuquerque, will water freeze at a lower temperature or a higher temperature than at 1 atmosphere? higher Will water boil at a higher or lower temperature, than at 1 atmosphere? lower



Part C – Phase Diagram for Carbon Dioxide.

14. At 1 atmosphere and room temperature (25°C), would you expect solid carbon dioxide to melt to the liquid phase, or sublime to the gas phase? sublime to gas
15. Some industrial processes require carbon dioxide. The carbon dioxide is stored on-site in large tanks as liquid carbon dioxide. Assuming we lived at sea level (1 atm), how could carbon dioxide be liquefied? Increase the pressure



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