

Lab Report 9: Synthesis of Isopentyl Acetate (Banana Oil)

Your Name: Jack Mladic Lab Partner: Danie Lodin Section: 3

All pages must be stapled together, with this being the top page.

Part 1: Experimental Data and Results (include appropriate unit) (2 pt)

1. Draw a **balanced** reaction equation showing the conversion of acetic acid and isopentyl alcohol to banana oil. Then fill out the table below.

	Acetic acid	Isopentyl alcohol
Mass used in lab	5.0 g	4.0 g
M.W.	60.052 g/mol	130.18 g/mol
Moles	.08326 moles	.03072 moles

2.

Theoretical yield of banana oil (show your calculation)	$5.0\text{g acetic acid}/(60.052\text{g/mol}) = 0.08326 \text{ mol acetic acid}$ which would give $0.08326 \text{ mol of isopentyl alcohol}$ $0.08326 \text{ mol of isopentyl alcohol} * 130.18\text{g/mol} = 10.839 \text{ g of banana oil}$
Actual yield	3.02g
% yield	$\text{Actual/theoretical} = 3.02/10.839 = 0.2786 = 28\%$

Part 2: Spectral Analysis of the product

2. (1pt) Carefully examine your IR spectrum for the product and compare it to the standard in the SDBS database (attached both yours and standard's). Did your IR indicate that all **isopentyl alcohol** had completely reacted? Why?

Isopentyl Acetate Lab Report

Ensheng Dong



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