

# Isotopes Practice Set

- What does "atomic number" mean?  $Z$  is the number of protons in the nucleus.
- What does the number next to isotopes mean? The mass number of the isotope.
- How do you find isotopes?  $A - Z = N$  (mass number - atomic number = number of neutrons)

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

  

	Magnesium-24	Magnesium-26
# of protons	12	12
# of neutrons	12	14
# of electrons	12	12

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

  

	Sulfur-32	Sulfur-34
# of protons	16	16
# of neutrons	16	18
# of electrons	16	16

Fill in the isotopes names including atomic numbers, including atomic numbers from the chart. Use your periodic table and the information provided.

	Isotope: $^{35}_{17}\text{Cl}$	Isotope: $^{37}_{17}\text{Cl}$
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Atom: $^{12}_{6}\text{C}$	Atom: $^{13}_{6}\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Carbon: $^{12}_{6}\text{C}$	Carbon: $^{13}_{6}\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Chlorine: $^{35}_{17}\text{Cl}$	Chlorine: $^{37}_{17}\text{Cl}$
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

# Isotope Practice Set

**ML Morrison**



**Isotope Practice Set:**

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