

1. Nitrogen and hydrogen react to form ammonia gas according to the following equation.



Z: 2  
I: 26

Z: 14  
I: 36

Molar mass

$N_2: 2 \times 14.0$

a. If 56.0 grams of nitrogen are used up by the reaction, how many grams of ammonia will be produced?

56.0g N <sub>2</sub>	1 mol N <sub>2</sub>	2 mol NH <sub>3</sub>	2 x 17.0g NH <sub>3</sub>
1	28.0g N <sub>2</sub>	1 mol N <sub>2</sub>	1 mol NH <sub>3</sub>
			<u>68.0g NH<sub>3</sub></u>

b. How many grams of hydrogen must react if the reaction needs to produce 68.0 grams of ammonia?

68.0g NH <sub>3</sub>	1 mol NH <sub>3</sub>	3 mol H <sub>2</sub>	3 x 2.0g H <sub>2</sub>
1	17.0g NH <sub>3</sub>	1 mol NH <sub>3</sub>	1 mol H <sub>2</sub>
			<u>6.0g H<sub>2</sub></u>

2. Aluminum metal reacts with zinc chloride to produce zinc metal and aluminum chloride.



Al: 27  
Zn: 65  
Cl: 35.5

a. A mass of 45.0 grams of aluminum will react with how many grams of zinc chloride?

45.0g Al	1 mol Al	3 mol ZnCl <sub>2</sub>	3 x 136.0g ZnCl <sub>2</sub>
1	27.0g Al	1 mol Al	1 mol ZnCl <sub>2</sub>
			<u>513.0g ZnCl<sub>2</sub></u>

b. What mass of aluminum chloride will be produced if 22.5 grams of zinc chloride are used up in the reaction?

22.5g ZnCl <sub>2</sub>	1 mol ZnCl <sub>2</sub>	2 mol AlCl <sub>3</sub>	2 x 133.5g AlCl <sub>3</sub>
1	136.0g ZnCl <sub>2</sub>	1 mol ZnCl <sub>2</sub>	1 mol AlCl <sub>3</sub>
			<u>74.7g AlCl<sub>3</sub></u>

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