

Entropy WS 11

Section A

- 1 The equation for the reaction of CO_2 and H_2O to produce glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, and O_2 is shown below.



Standard entropies are given in the table below.

Substance	$\text{CO}_2(\text{g})$	$\text{H}_2\text{O}(\text{l})$	$\text{O}_2(\text{g})$
$S^\circ / \text{J K}^{-1} \text{ mol}^{-1}$	214	70	205

- (a) (i) Calculate the standard entropy of glucose.

$$\begin{aligned} 256 &= [n + 6(205)] - [6(214) + 6(70)] \\ -256 &= (n + 1230) - (1284 + 420) \\ -256 &= n + 1230 - 1704 \\ n &= -256 + 474 \\ &= +218 \text{ J K}^{-1} \text{ mol}^{-1} \end{aligned}$$

$$S^\circ = \dots\dots\dots +218 \dots\dots\dots \text{ J K}^{-1} \text{ mol}^{-1} \quad [2]$$

- (ii) Calculate ΔG , in kJ mol^{-1} , at 25°C .

Show all your working.

$$\begin{aligned} \Delta G &= \Delta H - T\Delta S \\ &= +2879 - 298 \left(\frac{-256}{1000} \right) \\ &= 2879 + 76.28 \\ &= 2955.28 \end{aligned}$$

$$\Delta G = \dots\dots\dots +2955.28 \dots\dots\dots \text{ kJ mol}^{-1} \quad [2]$$

- (iii) Explain why this reaction is not feasible at any temperature.

As ΔH is endothermic and ΔS is negative, the value for ΔG will always be positive. Therefore this reaction will not be feasible at any temperature.

[1]

Entropy Section Review Answers

Sebastian Brünink



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