

Enzyme Action: Testing Catalase Activity

Many people are familiar with the hydrogen peroxide (H_2O_2) disinfectant. Hydrogen peroxide is used to clean cuts and abrasions. It is effective as an antiseptic and disinfectant. It is also used as a bleaching agent. Hydrogen peroxide is a colorless liquid with a strong, sharp odor. It is used in many household products. Hydrogen peroxide is a weak oxidizing agent. It is used in many industrial processes. Hydrogen peroxide is a colorless liquid with a strong, sharp odor. It is used in many household products. Hydrogen peroxide is a weak oxidizing agent. It is used in many industrial processes.

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In this experiment, you will measure the rate of oxygen evolution during the decomposition of hydrogen peroxide. The rate of oxygen evolution is measured by measuring the volume of oxygen gas evolved over time. The rate of oxygen evolution is measured by measuring the volume of oxygen gas evolved over time.

- measuring the rate of appearance of a product in the reaction.
- measuring the rate of disappearance of a reactant in the reaction.
- measuring the percentage of a reactant that has reacted.



In this experiment, you will measure the rate of oxygen evolution during the decomposition of hydrogen peroxide. The rate of oxygen evolution is measured by measuring the volume of oxygen gas evolved over time. The rate of oxygen evolution is measured by measuring the volume of oxygen gas evolved over time.

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CONCLUSIONS

In this experiment, you will:

- use an oxygen gas syringe to measure the percentage of hydrogen peroxide that has reacted.
- determine what factors affect the rate of oxygen evolution during the decomposition of hydrogen peroxide.

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Gunstream



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