

Use the table of standard reduction potentials from your textbook to complete the following table. The zinc-copper Daniell cell is given as an example.

Electrodes	Conc. of Each Cation	Half-Reactions	Cathode or Anode?	$E^\circ$ (table)	$E^\circ_{\text{cell}}$ (calc'd)	$E^\circ_{\text{cell}}$ (measured)
<b>A. Determine the <math>E^\circ</math> for the Standard Zn-Cu Galvanic Cell</b>						
Zn	1.0 M $\text{Zn}^{2+}$	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	+1.10 V	0.93 V
Cu	1.0 M $\text{Cu}^{2+}$	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
<b>B. Determine the <math>E^\circ</math> for the Cu-Zn Cell at Non-standard Concentrations</b>						
Zn	1.0 M $\text{Zn}^{2+}$	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.95 V
Cu	0.5 M $\text{Cu}^{2+}$	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
Zn	1.0 M $\text{Zn}^{2+}$	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.96 V
Cu	0.1 M $\text{Cu}^{2+}$	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
<b>C. Determine the <math>E^\circ</math> for Zn-Mg and Cu-Mg Galvanic Cells</b>						
Zn	1.0 M $\text{Zn}^{2+}$	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.82 V
Mg	1.0 M $\text{Mg}^{2+}$	$\text{Mg}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Mg}_{\text{s}}$	Cathode	-2.38 V		
Cu	1.0 M $\text{Cu}^{2+}$	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Anode	+0.34 V	_____ V	1.10 V
Mg	1.0 M $\text{Mg}^{2+}$	$\text{Mg}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Mg}_{\text{s}}$	Cathode	-2.38 V		

\*You must use the Nernst equation to calculate this  $E_{\text{cell}}$  voltage. (See Equation 8.)

# Galvanic Cell Experiment Report

**Rachel S Tattersall**



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