

### BUILD Vocabulary

**ionic bond** Chemical bond formed when one or more electrons are transferred from one atom to another.

**ion** atom that has a positive or negative charge.

**covalent bond** type of bond between atoms in which the electrons are shared.

**molecule** smallest unit of most compounds that displays all the properties of that compound.

**van der Waals forces** weak attraction that develops between oppositely charged regions of nearby molecules.

**Preface** The preface can be read or read "together." A compound is made up of elements joined together, and a covalent bond connects atoms that share electrons together. **How are covalent bonds different from ionic bonds?**

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## Chemical Bonds

**KEY QUESTION** What are the main types of chemical bonds?

The atoms in compounds are held together by chemical bonds. Bonds are formed by the electrons in the outer shell, which are called valence electrons. The main types of chemical bonds are ionic bonds and covalent bonds.

**Ionic Bonds** **ionic bonds** are formed when one or more electrons are transferred from one atom to another. This transfer causes the atom losing the electrons to become positively charged. The atom gaining the electrons becomes negatively charged. The positively and negatively charged atoms become **ions**. The oppositely charged ions are attracted to each other and form an ionic bond.

**Covalent Bonds** Sometimes atoms share electrons rather than transferring them. These shared electrons move around the nuclei of both atoms. The bond formed by sharing electrons is called a **covalent bond**. Atoms can share between two and six electrons. The structure that results is known as a **molecule**. A molecule is the smallest unit of most compounds.

**Weak Interactions** The strongest chemical bonds are ionic or covalent. Weak interactions between atoms are also important. Within living organisms, many molecules interact only briefly. **Van der Waals forces** are one type of weak interaction between molecules. They occur when molecules are very close together.

Hydrogen bonds are another type of weak interaction. They occur between a hydrogen atom of one molecule and an oxygen or nitrogen atom of a neighboring molecule.

### Visualizing with Covalent Bonds

1. What kind of chemical bond holds the water molecule together? \_\_\_\_\_
2. What particles are shared between the atoms in the molecule? \_\_\_\_\_
3. What is the other main type of chemical bond? \_\_\_\_\_
4. What is the difference between covalent and ionic bonds? \_\_\_\_\_



Water molecule (H<sub>2</sub>O)

**LEARNING LOOK** Cause and Effect As you read the part of the lesson under "The Water Molecule," use the table below to list the causes and effects of the properties of water. Many of the causes may be the same, but each characteristically will have a different effect. Part of the table is completed for you.

Characteristics of Water	Cause	Effect
Hydrogen Bonding		The slightly negative end of a water molecule is attracted to the slightly positive end of another water molecule.
Expands Upon Freezing		
Dissolves Many Substances	Water forms hydrogen bonds.	
Cohesion		
Adhesion		
Heat Capacity		

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**Kenneth R Miller, Joseph S Levine**



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## **Table of Contents Miller Levine Biology Study Guide Answers**

1. Understanding the eBook Miller Levine Biology Study Guide Answers
  - The Rise of Digital Reading Miller Levine Biology Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Miller Levine Biology Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Miller Levine Biology Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Miller Levine Biology Study Guide Answers
  - Personalized Recommendations
  - Miller Levine Biology Study Guide Answers User Reviews and Ratings
  - Miller Levine Biology Study Guide Answers and Bestseller Lists

5. Accessing Miller Levine Biology Study Guide Answers Free and Paid eBooks
  - Miller Levine Biology Study Guide Answers Public Domain eBooks
  - Miller Levine Biology Study Guide Answers eBook Subscription Services
  - Miller Levine Biology Study Guide Answers Budget-Friendly Options
6. Navigating Miller Levine Biology Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Miller Levine Biology Study Guide Answers Compatibility with Devices
  - Miller Levine Biology Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Miller Levine Biology Study Guide Answers
  - Highlighting and Note-Taking Miller Levine Biology Study Guide Answers
  - Interactive Elements Miller Levine Biology Study Guide Answers
8. Staying Engaged with Miller Levine Biology Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Miller Levine Biology Study Guide Answers
9. Balancing eBooks and Physical Books Miller Levine Biology Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Miller Levine Biology Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Miller Levine Biology Study Guide Answers
  - Setting Reading Goals Miller Levine Biology Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Miller Levine Biology Study Guide Answers
  - Fact-Checking eBook Content of Miller Levine Biology Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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