

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<b>58.5 g</b>

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<b>84.0</b>

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<b>159.6</b>

# Modeling Chemistry Unit 7 Test Answer

**John Bodenhan Taylor, D. J. Trigg**



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## **Table of Contents Modeling Chemistry Unit 7 Test Answer**

1. Understanding the eBook Modeling Chemistry Unit 7 Test Answer
  - The Rise of Digital Reading Modeling Chemistry Unit 7 Test Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Test Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 Test Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Test Answer
  - Personalized Recommendations
  - Modeling Chemistry Unit 7 Test Answer User Reviews and Ratings

- Modeling Chemistry Unit 7 Test Answer and Bestseller Lists
- 5. Accessing Modeling Chemistry Unit 7 Test Answer Free and Paid eBooks
  - Modeling Chemistry Unit 7 Test Answer Public Domain eBooks
  - Modeling Chemistry Unit 7 Test Answer eBook Subscription Services
  - Modeling Chemistry Unit 7 Test Answer Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 7 Test Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 7 Test Answer Compatibility with Devices
  - Modeling Chemistry Unit 7 Test Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Test Answer
  - Highlighting and Note-Taking Modeling Chemistry Unit 7 Test Answer
  - Interactive Elements Modeling Chemistry Unit 7 Test Answer
- 8. Staying Engaged with Modeling Chemistry Unit 7 Test Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Unit 7 Test Answer
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Test Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Test Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Test Answer
  - Setting Reading Goals Modeling Chemistry Unit 7 Test Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Test Answer
  - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Test Answer
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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