

## Guided Practice: Equilibrium

### Collision Theory & Reaction Rates

#### A - How Chemical Reactions Occur

1. Using the collision model, explain at a molecular level why reactions proceed **FASTER** when concentrations increase [4.5g]

#### B - Conditions that Affect Reaction Rates

2. What is activation energy? After consulting your text (Fig 17.2 and 17.3) draw a graph showing a reaction pathway - labeling the activation within the graph.



3. Use the collision theory to explain why reactions should occur more **SLOWER** at lower temperatures.
4. How does a catalyst speed up a reaction? (Refer to the activation energy labeled in #2 in your answer.)
5. Explain the difference between a homogeneous reaction and a heterogeneous reaction. Provide an example of each.
6. List the factors that affect reaction rates; explain how each factor can affect the rate of a reaction:
- a.
  - b.
  - c.
  - d.
7. After reading the definition for equilibrium, provide an example of equilibrium in everyday life.
8. Using your understanding of equilibrium, provide your own definition for chemical equilibrium.

# Equilibrium Unit Webquest And Notes

**M Mark**



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