

Electron Configuration Practice Worksheet

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 5s^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4 6d^1$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$ → $[Kr] 5s^1 4d^{10}$ (exception → more stable)
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$ (last electron in d of shell)

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ Sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ Rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ Antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ Osmium
- 15) $[Rn] 7s^2 5f^{11}$ Terbium

Determine which of the following electron configurations are not valid:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (3d not there)
- 17) $1s^2 2s^2 2p^6 3s^2 3d^5$ not valid (3d comes after 3p)
- 18) $[Ra] 7s^2 5f^6$ not valid [Ra] is not a noble gas
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ valid (iodine)
- 20) $[Xe]$ not valid (noble gases cannot be abbreviated)

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