

## Electron Configuration Practice Worksheet

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^9 6p^6 7s^2 5f^4$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^2 4d^9$
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  rubidium
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  antimony
- 14)  $[Xe] 6s^2 4f^{14} 5d^5$  osmium
- 15)  $[Rn] 7s^2 5f^{11}$  einsteinium

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{13} 4p^5$  not valid
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^3$  not valid
- 18)  $[Rn] 7s^2 5f^8$  not valid
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$  valid
- 20)  $[Xe]$  not valid (an element can't be its own configuration)

# Electron Configuration Ws Answer Key

**Victor M. Corman**



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## **Table of Contents Electron Configuration Ws Answer Key**

1. Understanding the eBook Electron Configuration Ws Answer Key
  - The Rise of Digital Reading Electron Configuration Ws Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Ws Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Ws Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Ws Answer Key
  - Personalized Recommendations
  - Electron Configuration Ws Answer Key User Reviews and Ratings
  - Electron Configuration Ws Answer Key and Bestseller Lists
5. Accessing Electron Configuration Ws Answer Key Free and Paid eBooks

- Electron Configuration Ws Answer Key Public Domain eBooks
  - Electron Configuration Ws Answer Key eBook Subscription Services
  - Electron Configuration Ws Answer Key Budget-Friendly Options
6. Navigating Electron Configuration Ws Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Electron Configuration Ws Answer Key Compatibility with Devices
  - Electron Configuration Ws Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Electron Configuration Ws Answer Key
  - Highlighting and Note-Taking Electron Configuration Ws Answer Key
  - Interactive Elements Electron Configuration Ws Answer Key
8. Staying Engaged with Electron Configuration Ws Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Ws Answer Key
9. Balancing eBooks and Physical Books Electron Configuration Ws Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Ws Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Ws Answer Key
- Setting Reading Goals Electron Configuration Ws Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Ws Answer Key
- Fact-Checking eBook Content of Electron Configuration Ws Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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