

8 • Electron Configurations & Periodicity

ANSWERS TO STUDY QUESTIONS

- Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
 - phosphorus [Ne] $3s^2 3p^3$
 - nickel [Ar] $3d^8 4s^2$
 - osmium [Xe] $4f^{14} 5d^6 6s^2$
 - californium [Rn] $5f^{10} 7s^2$
 - titanium [Ar] $3d^2 4s^2$
- Which orbital is filled following these orbitals?
 - 3d is followed by 4p
 - 4s is followed by 3d
 - 5p is followed by 6s
 - 5f is followed by 6d
- How many electrons can be accommodated in
 - a d subshell 10; two in each of 5 orbitals
 - a set of f orbitals 14; two in each of 7 orbitals
 - the $n = 4$ shell 32; $2s^2$
 - the 7s orbital 2
 - a p_x orbital? 2
- What is wrong with the following ground state electron configurations?

a. $\begin{array}{c} 3d \\ \boxed{} \boxed{} \boxed{} \boxed{} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{} \boxed{} \end{array}$	a. OK, 4s is filled before 3d.
b. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{1} \boxed{1} \end{array}$	b. 4s should be filled before 4p subshell.
c. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{1} \boxed{1} \end{array}$	c. 4s cannot have 2 electrons with same spin.
d. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{} \boxed{} \end{array}$	d. OK, stability of half-filled 3d subshell.
e. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{} \end{array}$	e. 4p should first fill one orbital at a time.

- How many unpaired electrons are there in
 - a nitrogen atom 3 unpaired electrons ($2s^2 2p^3$)
 - an iodine atom 1 unpaired electrons ($5s^2 5p^5$)
 - a nickel (II) cation 2 unpaired electrons ($3d^8$)
 - an oxide ion no unpaired electrons (closed shell [Ne])

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Zhe Xu



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