

1. In general, the atomic mass of elements increases as the atomic number increases. Find the pairs of elements in the periodic table that are exceptions to this generalization.

2. Write the symbol for the elements in the following locations.
- period 3, Group IIIA
 - period 1, Group VIIIA
 - period 4, Group IIB
 - period 6, Group VA

3. Name the third period element that has a half-filled *p* orbital.

4. Identify the Group (by group number and Family (if applicable)) for the following elements on the periodic table that have the following "outer electron configurations".

- s^2
- s^2p^5
- electrons filling the *d* orbitals
- s^2p^6

5. How many Valence electrons are in the following?
- elements of the oxygen family (same column)
 - Na atom
 - The element in Group IIIA, period 5
 - The element with the electron configuration $1s^22s^22p^5$

6. Predict the number of protons and electrons in stable IONS of radium and iodine.

7. What does the principle quantum number, *n*, designate?

8. Predict the common ions formed when atoms of the elements listed gain or lose electrons. Then name the noble gas with which the ion is isoelectronic.

- magnesium
- chlorine
- aluminum
- potassium
- sulfur
- barium
- phosphorus



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S Nieto



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