

Name: _____

Properties of Acids and Bases

Introduction: In an aqueous solution, an acid yields an H^+ ion. Bases yield an OH^- ion. But what other properties do they have? Acids react with metals in a single replacement reaction:



Also acids react with a carbonate:



A neutralization reaction can occur between an acid and a base to form salt and water. In this lab we will observe those reactions as well as test their acidity and pH.

Purpose: To determine the properties of acids and bases.

Pre-Lab:

1. Create a data table for results found in step one and two.
2. Identify HCl, $HC_2H_3O_2$, and NaOH as an acid or base and explain.
3. When is phenolphthalein colorless? When is it pink?
4. Write a balanced double replacement equation for the reaction of HCl and $CaCO_3$.
5. Write a balance reaction for the reaction of HCl and NaOH.
6. What type of reaction occurs between HCl and NaOH?

Materials: test tubes and rack, pipettes, spatula, spot plate, splint, 6M HCl, 6M $HC_2H_3O_2$, 0.5M NaOH, 1M NaOH, phenolphthalein, litmus paper, pH paper, zinc, magnesium, copper, $CaCO_3$.

Procedure:

1. Test each solution at this lab station with red and blue litmus paper and pH paper. Then decide if they are acids or bases. Record the substance name and results.

Substance	Red litmus	Blue litmus	pH paper	Acid or Base?

Lab 20 Properties Of Acids And Bases

Thomas Griffiths



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Table of Contents Lab 20 Properties Of Acids And Bases

1. Understanding the eBook Lab 20 Properties Of Acids And Bases
 - The Rise of Digital Reading Lab 20 Properties Of Acids And Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 20 Properties Of Acids And Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 20 Properties Of Acids And Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 20 Properties Of Acids And Bases
 - Personalized Recommendations
 - Lab 20 Properties Of Acids And Bases User Reviews and Ratings

- Lab 20 Properties Of Acids And Bases and Bestseller Lists
- 5. Accessing Lab 20 Properties Of Acids And Bases Free and Paid eBooks
 - Lab 20 Properties Of Acids And Bases Public Domain eBooks
 - Lab 20 Properties Of Acids And Bases eBook Subscription Services
 - Lab 20 Properties Of Acids And Bases Budget-Friendly Options
- 6. Navigating Lab 20 Properties Of Acids And Bases eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab 20 Properties Of Acids And Bases Compatibility with Devices
 - Lab 20 Properties Of Acids And Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 20 Properties Of Acids And Bases
 - Highlighting and Note-Taking Lab 20 Properties Of Acids And Bases
 - Interactive Elements Lab 20 Properties Of Acids And Bases
- 8. Staying Engaged with Lab 20 Properties Of Acids And Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 20 Properties Of Acids And Bases
- 9. Balancing eBooks and Physical Books Lab 20 Properties Of Acids And Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 20 Properties Of Acids And Bases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab 20 Properties Of Acids And Bases
 - Setting Reading Goals Lab 20 Properties Of Acids And Bases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lab 20 Properties Of Acids And Bases
 - Fact-Checking eBook Content of Lab 20 Properties Of Acids And Bases
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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