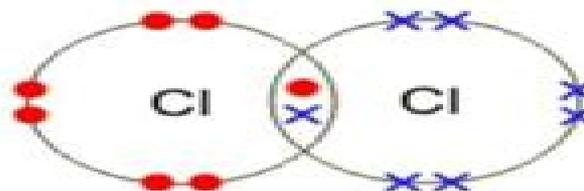


Covalent bonding

A non-metal metal and another non-metal can bond together by **sharing** electrons so that both electrons can achieve a **full outer shell**. This is called **covalent bonding**. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared

1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
2. Write the symbols of the atoms involved in the middle of each circle
3. Draw the **electrons** for one atom as dots (o) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the **overlapping** section.
5. Count up the electrons for **EACH** atom and check that they each now have a **full outer shell** (2 electrons for the first shell and 8 for the next 2 shells)

TASK: Draw dot and cross diagrams for the following covalent compounds.

F_2		CH_4	
HCl		CO_2	
H_2S		H_2O	
HF		NH_3	

Challenge - O_2

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**Mark Herbert McAdon, William A.
Goddard, California Institute of
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Table of Contents Metallic Bonding Worksheet Answers

1. Understanding the eBook Metallic Bonding Worksheet Answers
 - The Rise of Digital Reading Metallic Bonding Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Metallic Bonding Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metallic Bonding Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metallic Bonding Worksheet Answers
 - Personalized Recommendations
 - Metallic Bonding Worksheet Answers User Reviews and Ratings
 - Metallic Bonding Worksheet Answers and Bestseller Lists
5. Accessing Metallic Bonding Worksheet Answers Free and Paid eBooks
 - Metallic Bonding Worksheet Answers Public Domain eBooks
 - Metallic Bonding Worksheet Answers eBook Subscription Services
 - Metallic Bonding Worksheet Answers Budget-Friendly Options
6. Navigating Metallic Bonding Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Metallic Bonding Worksheet Answers Compatibility with Devices
 - Metallic Bonding Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Metallic Bonding Worksheet Answers
 - Highlighting and Note-Taking Metallic Bonding Worksheet Answers
 - Interactive Elements Metallic Bonding Worksheet Answers
8. Staying Engaged with Metallic Bonding Worksheet Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Metallic Bonding Worksheet Answers
9. Balancing eBooks and Physical Books Metallic Bonding Worksheet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Metallic Bonding Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Metallic Bonding Worksheet Answers
- Setting Reading Goals Metallic Bonding Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Metallic Bonding Worksheet Answers
- Fact-Checking eBook Content of Metallic Bonding Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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