

## Electron Configuration Practice Worksheet

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4 6d^1$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^1 4d^{10}$  →  $[Kr] 5s^1 4d^{10}$  (exception → more stable)
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$  ← last electron in of shell

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  Sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  Rubidium
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  Antimony
- 14)  $[Xe] 6s^2 4f^{14} 5d^6$  Osmium
- 15)  $[Rn] 7s^2 5f^{11}$  Terbium

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (3d not there)
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$  not valid (3d comes after 3p)
- 18)  $[Ra] 7s^2 5f^6$  not valid [Ra] is not a noble gas
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$  valid (iodine)
- 20)  $[Xe]$  not valid (noble gases cannot be abbreviated)

# Electron Configuration Practice Packet Key

**G Thomas**



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