

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Objectives: (1) (10-1) P. 41 - Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are associated with experimental evidence. (10-1) P. 42 - Explain the relative physical and chemical behavior of atoms by describing the structure of atoms in terms of positive nucleus, orbiting electrons, and the periodic table. (10-1) P. 43 - Explain properties of atoms and their changes in the periodic table in terms of patterns of their properties related to P. 43, as well as the structure of atoms in the periodic table (atomic structure, periodicity, shell model, Bohr model).

ANSWER KEY

Directions: Match the atomic theory element to its correct location.

Element	Location	Atomic Theory Element
1. C	Center (Nucleus)	1. Everything is made of atoms (atoms are the smallest, building up everything)
2. e	Circle (Orbit)	2. Matter (Building Block) consists mostly of positively charged nuclei and negatively charged electrons with neutral electrical charge (not neutral)
3. p	Proton(s) (+)	3. Atoms are indivisible (cannot be created)
4. n	Neutron(s) (0)	4. Group Number = Valency Element
5. M	Model (Bohr)	5. Electrons flow
6. E	Electrons (-)	6. Matter is made of tiny particles (atoms, molecules, ions)

Directions: Use the periodic table to answer the following questions:

1. **100** = **P. 43** - The number of protons equals the number of electrons.
2. **atomic mass** minus the **atomic number** will equal the **number of neutrons**.
3. **P. 43** = **100** - The number of protons in an atom (neutral atom) is the same as the number of electrons (negative).

Electrons In Atoms Answer Key Study Guide

Douglas Brandt



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