

Intro to Energy, Enzymes, and Metabolism

Sept 24th

Concepts in Energy and Thermodynamics

- Energy
- Reactions
- Entropy
- 2nd law of thermodynamics
- Free energy
- Kinetics of chemical reactions
- ATP
- Catalysis and Enzymes

Energy is the ability to do work, and it will determine the **direction** and **rate** of reaction ($aA + bB \leftrightarrow cC + dD$)

Enthalpy (H)

- The total energy of a system

Free Energy (G)

- The amount of a systems energy that is available and can be used to promote change or do

Entropy (S)

- The degree of disorder of a system
- Higher Entropy \rightarrow More Disordered
- Lower Entropy \rightarrow Highly Ordered

T \rightarrow Temperature

Total Energy = Usable Energy + Unusable Energy

$$(H = G + TS)$$

$$(\Delta G = \Delta H - T \Delta S)$$

- Negative Delta G \rightarrow loss of energy = Spontaneous reaction

2nd Law of Thermodynamics

- The state of a natural system itself can be reserved, but **not** without increasing the entropy of the systems surroundings
- The change in free energy determines reaction direction
- Spontaneous doesn't need a input of energy
- Delta G will tell you what direction it will go, not the rate of the reaction

Mastering Biology Ch6

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