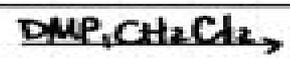
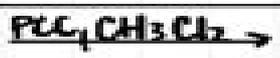


REAGENTS	NAME OF REACTION	WHAT IS PRODUCED	NOTES	FOLLOW THE MECH?
Mild oxidizing agents (see list)	Oxidation of Primary Alcohols (anhydrous)	$R-CH_2OH \rightarrow R-CHO$ 1° alcohol → aldehyde	• Must use a mild oxidizing agent • Harsh oxidizing agents proceed to carboxylic acid	NO (Ch. 12.10)
1) O ₃ 2) DMS	Ozonolysis of Alkenes	$R_2C=CR_2 \rightarrow R_2C=O + R_2C=O$ di-sub alkene → aldehyde	• Can also be used to form ketone when tetrasubstituted	NO (Ch. 8.12)
1) R ₂ B-H 2) H ₂ O ₂ , NaOH	Hydroboration-oxidation of Terminal Alkyne	$R-C\equiv CH \rightarrow R-CH_2-CHO$ Terminal alkyne → aldehyde	Anti-Markovnikov addition of H ₂ O across π-bond, then tautomerization	NO (Ch. 9.7)
Mild or Harsh oxidizing (see list)	Oxidation of Secondary Alcohols	$R-CH(OH)-R \rightarrow R-C(=O)-R$ 2° alcohol → ketone	• Mild or Harsh oxidizing work. • Won't oxidize past ketone	NO (Ch. 12.1)

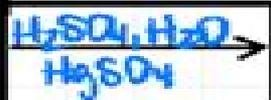
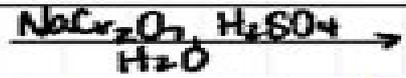
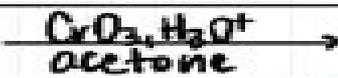
Oxidizing agents:

- 1° alcohol → aldehyde (mild stops) → carboxylic acid (harsh)
- 2° alcohol → ketone (mild or harsh)

Mild:



Harsh:



Acid-catalyzed Hydration of Terminal Alkyne



Markovnikov addition of H₂O across π-bond, then tautomerization

NO (Ch. 9.7)



Friedel-Crafts Acylation



Will not react if too deactivated. Needs a Lewis acid present.

YES (Ch. 18.6)

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R Bogdan

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1. Understanding the eBook Mastering Chemistry Ch 19 Key
 - The Rise of Digital Reading Mastering Chemistry Ch 19 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Ch 19 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Ch 19 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Ch 19 Key
 - Personalized Recommendations
 - Mastering Chemistry Ch 19 Key User Reviews and Ratings
 - Mastering Chemistry Ch 19 Key and Bestseller Lists
5. Accessing Mastering Chemistry Ch 19 Key Free and Paid eBooks

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 - Mastering Chemistry Ch 19 Key eBook Subscription Services
 - Mastering Chemistry Ch 19 Key Budget-Friendly Options
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 - Mastering Chemistry Ch 19 Key Compatibility with Devices
 - Mastering Chemistry Ch 19 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Mastering Chemistry Ch 19 Key
 - Highlighting and Note-Taking Mastering Chemistry Ch 19 Key
 - Interactive Elements Mastering Chemistry Ch 19 Key
8. Staying Engaged with Mastering Chemistry Ch 19 Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Ch 19 Key
9. Balancing eBooks and Physical Books Mastering Chemistry Ch 19 Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Ch 19 Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Ch 19 Key
- Setting Reading Goals Mastering Chemistry Ch 19 Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Ch 19 Key
- Fact-Checking eBook Content of Mastering Chemistry Ch 19 Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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