

Answer Key: Understanding pH

What is a log?

$$\text{Log } 10 = 1$$

$$\text{Log } 1 = 0$$

$$\text{Log } 100,000,000 = 8$$

$$\text{Log } 0.001 = -3$$

$$\text{Log } 1.0 \times 10^{-8} = -8$$

Explain why the following problem is more complicated than the previous ones.

Log 650 = **not a whole-number power of ten**

The number 650 is between 100 and 1000, so the answer must be between 2 and 3.

Using your Calculator

$$\text{Log } 650 = 2.812913357$$

Directions on how to enter this into my calculator: **Log button followed by 650 = (no need to use end parentheses).**

Calculating pH

$$\text{pH} = -\log [\text{H}^+], \text{ where brackets mean concentration in M}$$

Find pH when $[\text{H}^+] = 2.6 \times 10^{-6} \text{ M}$ **pH = 5.585026652 (for now, just report to lots of significant figures)**

Find pH when $[\text{H}^+] = 1.8 \times 10^{-6} \text{ M}$ **pH = 5.744727495**

Find pH when $[\text{H}^+] = 3.9 \times 10^{-4} \text{ M}$ **pH = 3.408935393**

Notice that it is easier to compare pH values since concentration values are so small.

The lower the $[\text{H}^+]$, the higher the pH. (or vice versa)

Note that pH values do not have units.

Acidic or Basic?

When $[\text{H}^+] > [\text{OH}^-]$, the solution is acidic

When $[\text{H}^+] < [\text{OH}^-]$, the solution is basic

When $[\text{H}^+] = [\text{OH}^-]$, the solution is neutral

Find $[\text{H}^+]$ in a neutral solution using the equation for K_w :

$$K_w = [\text{H}^+][\text{OH}^-] \text{ let } x = [\text{H}^+], \text{ since } [\text{H}^+] = [\text{OH}^-], K_w = x^2 = 1.0 \times 10^{-14} \text{ M}^2 \quad x = 1.0 \times 10^{-7} \text{ M} = [\text{H}^+]$$

What is the pH of this neutral solution?

$$\text{pH} = 7$$

Therefore, when $\text{pH} < \underline{7}$, the solution is acidic

When $\text{pH} > \underline{7}$, the solution is basic

When $\text{pH} = \underline{7}$, the solution is neutral

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F Rizvi



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