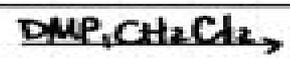
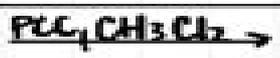


REAGENTS	NAME OF REACTION	WHAT IS PRODUCED	NOTES	FOLLOW THE MECH?
Mild oxidizing agents (see list)	Oxidation of Primary Alcohols (anhydrous)	$R-CH_2OH \rightarrow R-CHO$ 1° alcohol \rightarrow aldehyde	• Must use a mild oxidizing agent • Harsh oxidizing agents proceed to carboxylic acid	NO (Ch. 12.10)
1) O_3 2) DMS	Ozonolysis of Alkenes	$R_2C=CHR \rightarrow R_2C=O + R-CHO$ di-sub alkene \rightarrow aldehyde	• Can also be used to form ketone when tetrasubstituted	NO (Ch. 8.12)
1) R_2B-H 2) $H_2O_2, NaOH$	Hydroboration-oxidation of Terminal Alkyne	$R-C\equiv CH \rightarrow R-CH_2-CHO$ Terminal alkyne \rightarrow aldehyde	Anti-Markovnikov addition of H_2O across π -bond, then tautomerization	NO (Ch. 9.7)
Mild or Harsh oxidizing (see list)	Oxidation of Secondary Alcohols	$R-CH(OH)-R \rightarrow R-C(=O)-R$ 2° alcohol \rightarrow ketone	• Mild or Harsh oxidizing work. • Won't oxidize past ketone	NO (Ch. 12.1)

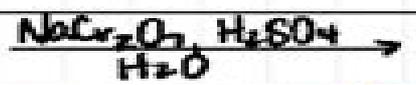
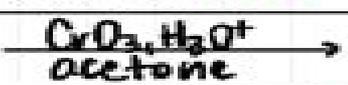
Oxidizing agents:

- 1° alcohol \rightarrow aldehyde (mild stops) \rightarrow carboxylic acid (harsh)
- 2° alcohol \rightarrow ketone (mild or harsh)

Mild:



Harsh:



H_2SO_4, H_2O $HgSO_4$	Acid-catalyzed Hydration of Terminal Alkyne	$R-C\equiv CH \rightarrow R-C(=O)CH_3$ Terminal Alkyne \rightarrow Ketone	Markovnikov addition of H_2O across π -bond, then tautomerization	NO (Ch. 9.7)
$Cl-C(=O)-R$ $AlCl_3$	Friedel-Crafts Acylation	$C_6H_6 \rightarrow C_6H_5-C(=O)R$ Benzene \rightarrow Benzoylic ketone	Will not react if too deactivated. Needs a Lewis acid present.	YES (Ch. 18.6)

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