

*Matter and Change, pp. 1-4*

- |    |    |   |
|----|----|---|
| 2. | 5. | c |
| 3. | 5. | c |
| 4. | 7. | c |
| 9. | 3. | a |
10. mass and volume
  11. group or family
  12. homogeneous
  13. products
  14. metal
  15. solid state
  16. chemical properties
  17. atom
  18. compound
  19. nonmetal
  20. mixture
  21. pure substance
  22. basic
  23. metalloid
  24. reactants: carbon and oxygen; product: carbon dioxide
  25. reactant: mercury(II) oxide; products: mercury and oxygen
  26. f
  27. e
  28. c
  29. a
  30. b
  31. i
  32. d
  33. h
  34. e
  35. physical change
  36. chemical change
  37. physical change
  38. physical change
  39. chemical change
  40. physical change
  41. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
  42. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
  43. Although energy can be absorbed or released in a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

*Measurements and Calculations, pp. 5-8*

- |      |      |
|------|------|
| a    | 5. c |
| 2. c | 6. c |
| 3. c | 7. a |
| 4. a | 8. a |
9. time
  10. mass
  11. density
  12. energy
  13. length or distance
  14. volume
  15. area
  16. qualitative
  17. quantitative
  18. qualitative
  19. quantitative
  20.  $3.00 \times 10^7$  km/s
  21. three
  22. 0.026 g or  $2.6 \times 10^{-2}$  g
  23.  $2.5 \times 10^{-4}$  L
  24. quantity
  25. derived units
  26. conversion factor
  27. 0.0432 kg or  $4.32 \times 10^{-2}$  kg
  28. 5400 mL or  $5.4 \times 10^3$  mL
  29. 300. K
  30.  $1.05 \times 10^3$  J
  31.  $3.51 \times 10^3$  cm
  32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
  33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
  34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
  35. The graph of two quantities that are directly proportional is a straight line through the origin.

# Modern Chemistry Answer Key Chapter 15

**Frank L. Wiseman**



## **Modern Chemistry Answer Key Chapter 15:**

*Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      *Principles of Modern Chemistry* David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom      *Modern Physical Chemistry* G.H. Duffey, 2013-11-11 In this new textbook on physical chemistry fundamentals are introduced simply yet in more depth than is common Topics are arranged in a progressive pattern with simpler theory early and more complicated theory later General principles are induced from key experimental results Some mathematical background is supplied where it would be helpful Each chapter includes worked out examples and numerous references Extensive problems review and discussion questions are included for each chapter More detail than is common is devoted to the nature of work and heat and how they differ Introductory Caratheodory theory and the standard integrating factor for  $dG_{rev}$  are carefully developed The fundamental role played by uncertainty and symmetry in quantum mechanics is emphasized In chemical kinetics various methods for determined rate laws are presented The key mechanisms are detailed Considerable statistical mechanics and reaction rate theory are then surveyed Professor Duffey has given us a most readable easily followed text in physical chemistry

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