

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Lei Shi



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Table of Contents Modern Chemistry Homework 15 Answers

1. Understanding the eBook Modern Chemistry Homework 15 Answers
 - The Rise of Digital Reading Modern Chemistry Homework 15 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Homework 15 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Modern Chemistry Homework 15 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Homework 15 Answers
 - Personalized Recommendations
 - Modern Chemistry Homework 15 Answers User Reviews and Ratings
 - Modern Chemistry Homework 15 Answers and Bestseller Lists
5. Accessing Modern Chemistry Homework 15 Answers Free and Paid eBooks
 - Modern Chemistry Homework 15 Answers Public Domain eBooks
 - Modern Chemistry Homework 15 Answers eBook Subscription Services
 - Modern Chemistry Homework 15 Answers Budget-Friendly Options

6. Navigating Modern Chemistry Homework 15 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Homework 15 Answers Compatibility with Devices
 - Modern Chemistry Homework 15 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Homework 15 Answers
 - Highlighting and Note-Taking Modern Chemistry Homework 15 Answers
 - Interactive Elements Modern Chemistry Homework 15 Answers
8. Staying Engaged with Modern Chemistry Homework 15 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Homework 15 Answers
9. Balancing eBooks and Physical Books Modern Chemistry Homework 15 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Homework 15 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Homework 15 Answers
 - Setting Reading Goals Modern Chemistry Homework 15 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Homework 15 Answers
 - Fact-Checking eBook Content of Modern Chemistry Homework 15 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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