

## Molarity Practice Worksheet

*Find the molarity of the following solutions:*

- 1) 0.5 moles of sodium chloride is dissolved to make 0.05 liters of solution.
  
- 2) 0.5 grams of sodium chloride is dissolved to make 0.05 liters of solution.
  
- 3) 0.5 grams of sodium chloride is dissolved to make 0.05 mL of solution.
  
- 4) 734 grams of lithium sulfate are dissolved to make 2500 mL of solution.
  
- 5)  $6.7 \times 10^{-2}$  grams of  $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$  are dissolved to make 3.5 mL of solution.
  
- 6) I have two solutions. In the first solution, 1.0 moles of sodium chloride is dissolved to make 1.0 liters of solution. In the second one, 1.0 moles of sodium chloride is added to 1.0 liters of water. Is the molarity of each solution the same? Explain your answer.

# Molarity Practice Unit Lesson 2

**D Kirk**

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Predicting Solubility of New Drugs Alex Avdeef, 2024-05-27 In pharmaceutical research

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**Loose-Leaf Version for Chemical Principles** Peter Atkins, Loretta Jones, Leroy Laverman, 2016-01-07 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do It also offers an exceptional level of support to help students develop their mathematical and problem solving skills For the new edition Chemical Principles now takes a modular approach with coverage organized as a series of brief Topics within 13 major areas of focus

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