

Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

1. What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

2. Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na: } & 22.990 \times 1 = 22.990 \\ \text{Cl: } & 35.453 \times 1 = 35.453 \end{aligned} \quad \Rightarrow \quad 58.443 \text{ g/mol}$$

3. What is the molar mass of NaOH?

$$\begin{aligned} \text{Na: } & 1 \times 22.990 = 22.990 \\ \text{O: } & 1 \times 15.999 = 15.999 \\ \text{H: } & 1 \times 1.008 = 1.008 \end{aligned} \quad \Rightarrow \quad 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

1. Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

2. Determine the number of moles in 80.00 g O₂.

$$0.2 \times 31.998 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 2.500 \text{ mol O}_2$$

3. Determine the volume of 0.24 mol of NO₂.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

4. Determine the number of moles in 7.5×10^{19} formula units of KNO₃.

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ mol KNO}_3$$

5. Determine the number of moles in 202.22g KNO₃.

$$\begin{aligned} \text{K: } & 1 \times 39.098 = 39.098 \\ \text{N: } & 1 \times 14.007 = 14.007 \\ \text{O: } & 3 \times 15.999 = 47.997 \end{aligned} \quad \Rightarrow \quad 101.102 \text{ g/mol} \quad \Rightarrow \quad 1 \text{ mol} = 101.102 \text{ g}$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

6. Determine the number of moles in 32.5L of N₂ at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

7. Determine the number of moles in 3.4×10^{24} formula units of NaCl.

$$3.4 \times 10^{24} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ mol NaCl}$$

8. Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

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