

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?
* Multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product * mole of reactant.

2. The reaction below produces ammonia for use in fertilizers.
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4. The reaction below is used to create soda for drinking. The reactants are hydrogen and oxygen.
 $2H_2 + O_2 \rightarrow 2H_2O$
If you start with 10 grams of hydrogen and 20 grams of oxygen, determine the limiting reactant.

5. Calculate the grams of hydrogen chloride produced.

6. How many grams of the excess reactant is present in a left over?

5. The reaction below is used to create soda for drinking. The reactants are hydrogen and oxygen.
 $2H_2 + O_2 \rightarrow 2H_2O$
If you start with 10 grams of hydrogen and 20 grams of oxygen, determine the limiting reactant.

6. How many grams of the excess reactant is present?

7. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

8. Calculate the grams of ammonia produced.

9. How many grams of the excess reactant is present in a left over?
* Multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

10. Calculate moles of product * mole of reactant.

ANSWER KEY INCLUDED

Limiting Reactant Problems Answer Key

Fred M. Dewey



Limiting Reactant Problems Answer Key:

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a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

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Understanding Chemistry Fred M. Dewey,1994 **Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook** McGraw Hill,2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change eBook: General, Organic and Biological Chemistry 2e SMITH,2012-02-16 eBook General Organic and Biological Chemistry 2e **SourceBook Version 2.1** ,1998 **Holt Chemistry** Ralph Thomas Myers,2004 **Student Solutions Manual: Ssm Chemistry** Deborah Wiegand,2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems *General Chemistry Key Concise Review Guide Sixth Edition* Darrell D. Ebbing,1999 The Science Teacher ,1992 Some issues are accompanied by a CD ROM on a selected topic *Preparatory Chemistry* Howard Stephen Stoker,1990 *Illinois Chemistry Teacher* ,1992 *Introduction to Chemical Principles* Howard Stephen Stoker,1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed preface **Basic Chemistry** Steven S Zumdahl,1996 **Chemistry** Charles E. Mortimer,1986

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