

# Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

## Chemistry

**LIMITING REACTANTS**

1. The reaction below produces ammonia for use in fertilizers.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If you start with 28 grams of  $N_2$  and 12 grams of  $H_2$ , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?

4. Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

5. Subtract mass of excess  $\times$  mole of product.

1. Write a balanced chemical equation for the reaction below and include the physical states of the reactants and products.

$$2H_2O_2(l) \rightarrow 2H_2O(l) + O_2(g)$$

If you have 100 grams of  $H_2O_2$  and 10 grams of  $O_2$ , determine the limiting reactant.

2. How many grams of water produced are produced?



1. Magnesium chloride is used to create salts for deicing roads. The reaction below produces magnesium chloride.

$$Mg + 2HCl \rightarrow MgCl_2 + H_2$$

If you start with 20 grams of magnesium and 20 grams of hydrochloric acid, determine the limiting reactant.

2. Calculate the grams of magnesium chloride produced.

3. How many grams of the excess reactant is present in a left over?

1. Magnesium chloride is used to create salts for deicing roads. The reaction below produces magnesium chloride.

$$Mg + 2HCl \rightarrow MgCl_2 + H_2$$

If you have 20 grams of magnesium and 20 grams of  $HCl$ , determine the limiting reactant.

2. How many grams of  $MgCl_2$  is produced?

1. The reaction below produces ammonia for use in fertilizers and a common household product. Determine the limiting reactant.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If you have 28 grams of  $N_2$  and 12 grams of  $H_2$ , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?

4. Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

5. Subtract mass of excess  $\times$  mole of product.

# ANSWER KEY INCLUDED

# Limiting Reactant Answer Key

**RM Cervero**



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