



# ISOTOPES PRACTICE SET



1. Write the symbol for the following isotopes:

a. 8 protons and 8 neutrons

c. atomic number is 11 and mass number is 23

b. 28 protons and 30 neutrons

d. 92 protons and mass number is 238

2. Fill in the table for the two atoms shown below.



	Atom 1	Atom 2
Number of Protons		
Number of Neutrons		
Mass Number		

3. Fill in the following table.

Isotope Symbol	${}^{40}_{19}\text{K}$	${}^{18}_9\text{F}$	
Atomic Number			16
Mass Number			
Number of Protons			
Number of Neutrons			15

Name : \_\_\_\_\_

Date : \_\_\_\_\_

# Isotopes Practice Set

**J Ma**



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