

Gravimetric Analysis of Metal Carbonate

Experiment Overview

The purpose of this lab is to determine the identity of a Group 1 metal carbonate compound by gravimetric analysis. The unknown is weighed and dissolved in water. A solution of calcium chloride is added to the metal carbonate solution to precipitate the carbonate ions as calcium carbonate. The precipitate is filtered, dried, and weighed. From the data, the formula weight and identity of the unknown metal carbonate is determined. Concepts covered in this lab are double replacement reaction (precipitate formation) and gravimetric analysis.

Background

The identity of group 1 metal M is determined by analyzing an unknown Group 1 metal carbonate, M_2CO_3 .

There are 3 main reactions in this lab:

1. Equation 1: $M_2CO_3 (s) \rightarrow 2M^+ (aq) + CO_3^{2-} (aq)$
2. Equation 2: $Ca^{2+} (aq) + CO_3^{2-} (aq) \rightarrow CaCO_3 (s)$
3. Equation 3: $CaCl_2 (aq) + M_2CO_3 (aq) \rightarrow CaCO_3 (s) + 2MCl (aq)$

The precipitated calcium carbonate is then filtered, dried, and weighed. *The moles of calcium carbonate, $CaCO_3$, are equal to the moles of Group 1 metal carbonate, M_2CO_3 , added to the original solution.* Dividing the mass of the unknown carbonate by the moles of calcium carbonate yields the formula weight, and thus the identity, of the Group 1 metal carbonate.

Pre-Lab Questions: Use the sample data table to answer the following questions:

Sample Data Table:

| | |
|---------------------------------|----------|
| Mass of crucible + M_2CO_3 | 12.627 g |
| Mass of crucible | 10.655 g |
| Mass of M_2CO_3 | 1.972 g |
| Mass of filter paper + $CaCO_3$ | 2.436 g |
| Mass of the filter paper | 0.598 g |
| Mass of the $CaCO_3$ | 1.838 g |
| Moles of $CaCO_3$ | mol |
| Molar mass of M_2CO_3 | g/mol |
| Identity of M_2CO_3 | |
| Percent error | % |

1. From the mass of $CaCO_3$, calculate the moles of $CaCO_3$ precipitated.
2. Calculate the molar mass of the unknown.
3. Calculate the molar mass of the following Group 1 metal carbonates:
 - a. Li_2CO_3
 - b. Na_2CO_3
 - c. K_2CO_3
4. What is the identity of M_2CO_3 ?
5. Calculate the percent error in the molar mass determination of M_2CO_3 , by comparing the experimentally determined molar mass of M_2CO_3 to the known molar mass of the appropriate metal carbonate.

Gravimetric Analysis Lab Answers

Alice S. Cohen



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English in Analytical Chemistry. Communicating about Methods & Techniques. Книга для студента Надежда Зорина, Александра Соболева, 2022-10-19 The Merck Report, 1899

Chemical Laboratory Mohamed Elzagheid, 2026-02-16 This book covers techniques in the chemical laboratory and safety procedures that are crucial to making the laboratory a safe workplace The book is divided into two sections the 1st comprehensively covering safety protocols in a chemical laboratory and the 2nd detailing important techniques to master This book can be utilized by graduate students laboratory technicians and laboratory chemists New to this edition Chemical Hazard Classification and Labeling Personal Protective Equipment Spill Management and Decontamination Procedures Waste Disposal and Environmental Compliance Biological and Radiological Hazards in Chemistry Labs Emergency Preparedness and Response Plans Safe handling of Nanomaterials and Emerging Risks Developing a Lab Safety Culture in Academic Institution Automation Sensors and Digital Tools in the Lab Microscale and Semi Microscale Techniques

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Cooperative Chemistry Lab Manual Cooper,2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations

Instructors Manual to Lab Manual Ralph Petrucci,William Harwood,Geoffrey Herring,2001

Exploring Chemical

Analysis Daniel C. Harris,2005 Exploring Chemical Analysis teaches students how to understand analytical results and how to use quantitative manipulations preparing them for the problems they will encounter

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