

GRAVIMETRIC ANALYSIS OF A CHLORIDE SALT

Date submitted: 14/03/2025

By: ()

Lab section:

TA

Purpose: The purpose of this experience is to determine the quantity of the chloride content in an unknown soluble salt.

Theory:

The assignment in this lab was to find the concentration of chloride ion in the given unknown salt sample. Gravimetric analysis is a quantitative analytical technique where the amount of an analyte is determined by measuring the mass of a solid precipitate.

Precipitation Reaction of Chloride Ions with Silver Ions: The reaction between chloride ions and silver ions forms a solid precipitate of silver chloride:



So, $n(\text{AgCl})$ is equal to $n(\text{Cl})$

Solubility of AgCl (s) in water:

According to the lab manual section "Theory" AgCl (s) is very insoluble, which lead to the completion of the reaction.

$$K_{sp} = [\text{Ag}^+][\text{Cl}^-] = 1.6 \times 10^{-10}$$

Calculation for the amount of precipitate lost by washing:

Even though, AgCl is insoluble during the washing process some of the precipitate was lost.

Therefore: $K_{sp} = [\text{Ag}(\text{aq})^+] \cdot [\text{Cl}(\text{aq})^-] = 1.6 \times 10^{-10}$

$$[x] \cdot [x] = 1.6 \times 10^{-10}$$

$$x = 1.3 \times 10^{-5}$$

- x is equal to 1.3×10^{-5} moles per litre

Precipitation is performed in an acidic environment:

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PT Brinkman

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

Gravimetric Analysis Of Chloride Salt Lab Report:

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