

Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter. It is a state of balance between the forward and reverse reactions.

What happens if you change the conditions? **Le Chatelier's principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will adjust itself to counteract the change and restore a new equilibrium.

What happens if you change the concentration of a reactant? **Le Chatelier's principle** states that if the concentration of a reactant is increased, the equilibrium will shift to the right to consume the added reactant.

What happens if you change the temperature? **Le Chatelier's principle** states that if the temperature is increased, the equilibrium will shift in the endothermic direction.

What happens if you change the pressure? **Le Chatelier's principle** states that if the pressure is increased, the equilibrium will shift towards the side with fewer moles of gas.

What happens if you change the volume? **Le Chatelier's principle** states that if the volume is decreased, the equilibrium will shift towards the side with fewer moles of gas.

What happens if you change the catalyst? **Le Chatelier's principle** states that a catalyst does not affect the position of equilibrium, but it does speed up the rate at which equilibrium is reached.

What happens if you change the surface area? **Le Chatelier's principle** states that a catalyst does not affect the position of equilibrium, but it does speed up the rate at which equilibrium is reached.

What happens if you change the concentration of a product? **Le Chatelier's principle** states that if the concentration of a product is increased, the equilibrium will shift to the left to consume the added product.

What happens if you change the temperature of an exothermic reaction? **Le Chatelier's principle** states that if the temperature is increased, the equilibrium will shift to the left to consume the added heat.

What happens if you change the temperature of an endothermic reaction? **Le Chatelier's principle** states that if the temperature is increased, the equilibrium will shift to the right to consume the added heat.

What happens if you change the pressure of a reaction with equal moles of gas on both sides? **Le Chatelier's principle** states that if the pressure is increased, the equilibrium will not shift.

What happens if you change the pressure of a reaction with unequal moles of gas? **Le Chatelier's principle** states that if the pressure is increased, the equilibrium will shift towards the side with fewer moles of gas.

What happens if you change the volume of a reaction with unequal moles of gas? **Le Chatelier's principle** states that if the volume is decreased, the equilibrium will shift towards the side with fewer moles of gas.

What are some examples of chemical equilibrium? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter. It is a state of balance between the forward and reverse reactions.

Nate Quick Study Guide

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