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# 5 INTERPRETING GRAPHICS

Use with Section 5.3



Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

$c = 3.00 \times 10^8 \text{ m/s}$   
 $E = hf$   
 $f = \frac{c}{\lambda}$

Table 1

Transition	$E(\text{J})$	Frequency $f(\text{Hz})$	Wave Length $\lambda(\text{m})$	Type of Radiation
$n=5 \rightarrow n=4$	$2.06 \times 10^{-18}$	$4.02 \times 10^{13} \frac{1}{\text{s}}$	$7.46 \times 10^{-6} \text{ m}$	infrared
$n=5 \rightarrow n=3$	$7.57 \times 10^{-18}$	$1.4 \times 10^{14} \frac{1}{\text{s}}$	$2.15 \times 10^{-6} \text{ m}$	infrared
$n=5 \rightarrow n=2$	$1.82 \times 10^{-17}$	$2.75 \times 10^{14} \frac{1}{\text{s}}$	$1.09 \times 10^{-6} \text{ m}$	infrared
$n=5 \rightarrow n=1$	$4.84 \times 10^{-17}$	$7.30 \times 10^{14} \frac{1}{\text{s}}$	$4.11 \times 10^{-7} \text{ m}$	Visible
$n=4 \rightarrow n=3$	$2.12 \times 10^{-17}$	$3.10 \times 10^{15} \frac{1}{\text{s}}$	$9.38 \times 10^{-8} \text{ m}$	ultraviolet
$n=4 \rightarrow n=2$	$4.91 \times 10^{-17}$	$7.41 \times 10^{15} \frac{1}{\text{s}}$	$4.05 \times 10^{-8} \text{ m}$	infrared
$n=4 \rightarrow n=1$	$1.35 \times 10^{-16}$	$2.94 \times 10^{16} \frac{1}{\text{s}}$	$1.28 \times 10^{-8} \text{ m}$	infrared
$n=3 \rightarrow n=2$	$4.56 \times 10^{-17}$	$6.88 \times 10^{15} \frac{1}{\text{s}}$	$4.36 \times 10^{-8} \text{ m}$	Visible
$n=3 \rightarrow n=1$	$2.09 \times 10^{-16}$	$3.15 \times 10^{16} \frac{1}{\text{s}}$	$9.52 \times 10^{-9} \text{ m}$	ultraviolet
$n=2 \rightarrow n=1$	$1.66 \times 10^{-16}$	$1.60 \times 10^{16} \frac{1}{\text{s}}$	$1.88 \times 10^{-8} \text{ m}$	infrared
$n=2 \rightarrow n=1$	$4.09 \times 10^{-17}$	$6.17 \times 10^{15} \frac{1}{\text{s}}$	$4.86 \times 10^{-8} \text{ m}$	Visible
$n=2 \rightarrow n=1$	$2.04 \times 10^{-17}$	$3.08 \times 10^{15} \frac{1}{\text{s}}$	$9.74 \times 10^{-8} \text{ m}$	ultraviolet
$n=2 \rightarrow n=1$	$3.00 \times 10^{-17}$	$4.57 \times 10^{15} \frac{1}{\text{s}}$	$6.56 \times 10^{-8} \text{ m}$	Visible
$n=2 \rightarrow n=1$	$1.04 \times 10^{-16}$	$2.93 \times 10^{15} \frac{1}{\text{s}}$	$1.62 \times 10^{-7} \text{ m}$	ultraviolet
$n=2 \rightarrow n=1$	$1.64 \times 10^{-16}$	$2.47 \times 10^{15} \frac{1}{\text{s}}$	$1.21 \times 10^{-7} \text{ m}$	ultraviolet

$f = \frac{c}{\lambda}$   
 $\lambda = \frac{c}{f}$

# Interpreting Graphics Chemistry Answer Key 16

**B Lingard**

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