

Lab 5: Equilibrium

In the equilibrium lab, the equilibrium constant for the formation of iron(II) thiocyanate was spectroscopically determined. In a similar experiment, the formation constant of $\text{Co}(\text{SCN})_4^{2-}$ was determined by mixing solutions of $\text{Co}(\text{NO}_3)_2$ and KSCN and measuring the absorbance of the blue $\text{Co}(\text{SCN})_4^{2-}$ formed in the reaction.



(a) What is the equilibrium expression for the reaction under study?

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] \cdot (4 \times [\text{SCN}^{-}])}$

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] \cdot [\text{SCN}^{-}]^4}$

$K = \frac{[\text{Co}^{2+}] \cdot [\text{SCN}^{-}]^4}{[\text{Co}(\text{SCN})_4^{2-}]}$

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] + (4 \times [\text{SCN}^{-}])}$

(b) Given the data in the table, please complete the chart below. The molar absorptivity of $\text{Co}(\text{SCN})_4^{2-}$ is $1800 \text{ M}^{-1} \text{ cm}^{-1}$ at 610 nm. Assume the Beer's Law plot has a slope of 1800 , and a y-intercept of zero.

Experimental Results

Volume of $1.00 \times 10^{-2} \text{ M Co}^{2+}$ (mL)	Volume of $8.00 \times 10^{-2} \text{ M SCN}^{-}$ (mL)	Initial $[\text{Co}^{2+}]$ (M)	Initial $[\text{SCN}^{-}]$ (M)	Absorbance at 610 nm	Calculated Equilibrium $[\text{Co}(\text{SCN})_4^{2-}]$ (M)
4.00	6.00	<input type="text" value="0.00025"/>	<input type="text" value="0.00040"/>	0.115	<input type="text" value="0.00025"/>

(c) Complete the reaction table for these results. (All entries should be in molarity.)

	$\text{Co}^{2+}(\text{aq})$	$+ 4 \text{SCN}^{-}(\text{aq})$	$\rightleftharpoons \text{Co}(\text{SCN})_4^{2-}(\text{aq})$
[initial]	<input type="text" value="0.00025"/>	<input type="text" value="0.00040"/>	0
[Δ]	<input type="text" value="0.00025"/>	<input type="text" value="0.00040"/>	<input type="text" value="0.00025"/>
[final]	<input type="text" value="0.00025"/>	<input type="text" value="0.00040"/>	<input type="text" value="0.00025"/>

(d) Calculate the formation constant for the reaction. (Enter your answer to three significant figures.)

$K_f =$

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