



PARTICLE THEORY AND STATES OF MATTER

1 What is particle theory?

2 All of matter is made of particles. Explain in which state of matter the particles are most closely packed?

3 Describe the arrangement of particles in a solid. How do the particles move?

4 The particles in a gas have the most energy. Explain why.

5 How are particles arranged in a liquid?

11 Describe the arrangement of particles in a solid. How do the particles move?

10 Describe the arrangement of particles in a solid. How do the particles move?

9 What is the arrangement of particles in a solid?

8 Describe the arrangement of particles in a solid. How do the particles move?

7 How are particles arranged in a liquid?

Particle Theory

All matter is made of particles. Particles are very small and they are in constant motion. They are separated by spaces. The particles in a solid are packed closely together. They vibrate about their fixed positions. The particles in a liquid are not so closely packed. They can move past each other. The particles in a gas are far apart and they move rapidly in all directions.

States of Matter

Solid: Particles are packed closely together. They vibrate about their fixed positions. They have a definite shape and volume.

Liquid: Particles are not so closely packed. They can move past each other. They have a definite volume but no definite shape.

Gas: Particles are far apart and they move rapidly in all directions. They have no definite shape or volume.

Changes of State

Melting: The change from a solid to a liquid. The particles in a solid are packed closely together. As they gain energy, they vibrate more and more. At a certain point, they break free from their fixed positions and they can move past each other. This is the melting point.

Boiling: The change from a liquid to a gas. The particles in a liquid are not so closely packed. As they gain energy, they move faster and faster. At a certain point, they break free from their fixed positions and they can move rapidly in all directions. This is the boiling point.

Condensing: The change from a gas to a liquid. The particles in a gas are far apart and they move rapidly in all directions. As they lose energy, they move slower and slower. At a certain point, they come together and they can move past each other. This is the condensing point.

Freezing: The change from a liquid to a solid. The particles in a liquid are not so closely packed. As they lose energy, they vibrate less and less. At a certain point, they come together and they can vibrate about their fixed positions. This is the freezing point.

Diagram:

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    graph LR
      Solid --> Liquid
      Liquid --> Gas
      Gas --> Solid
      Gas --> Liquid
      Liquid --> Solid
  
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