

## ANSWER KEY

### Mixed Stoichiometry Problems



a). How many moles of  $\text{H}_2$  would be required to produce 5.0 moles of water?

5.0 moles water

b). What mass of  $\text{H}_2\text{O}$  is formed when  $\text{H}_2$  reacts with 384 g of  $\text{O}_2$ ?

432g  $\text{H}_2\text{O}$



a). Balance this equation

Look above.

b). What mass of  $\text{H}_2\text{SO}_4$  would be required to react with 0.75 mol of  $\text{NaOH}$ ?

36.8 = 37g  $\text{H}_2\text{SO}_4$

d). If 30.0 g of  $\text{H}_2\text{SO}_4$  reacts with sodium hydroxide, what mass of water is produced?

11.0g  $\text{NaOH}$



a). How many grams of  $\text{SnF}_2$  are produced from the reaction of 30.0 g of  $\text{HF}$ ? 117g  $\text{SnF}_2$

b) How many moles of  $\text{HF}$  will be needed to react with 48 grams of Tin, Sn?

0.81 moles  $\text{HF}$

4. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide,  $\text{LiOH}$ , according to the following chemical equation.



If 20.0 moles of  $\text{CO}_2$  is exhaled, the average amount exhaled by a person each day, how much (in grams)  $\text{Li}_2\text{CO}_3(\text{s})$  is produced?

148g  $\text{Li}_2\text{CO}_3$



a). Balance the equation.

See above.

b). How many moles of sodium are required to produce 5.0 moles of  $\text{NaOH}$ , sodium hydroxide?

5.0 moles  $\text{Na}$

c). How many moles of  $\text{H}_2$  are produced in the reaction with 5.0 grams of water?

0.14 moles  $\text{H}_2$

# Mixed Stoichiometry Problems Key

**Frank Schaffer Publications, Joan  
DiStasio**



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