

Mixed Mole Problems Worksheet

Solve the following problems. Be sure to show all your work!

1. Find the number of moles of silver ions (
- Ag^+
-) or
- Ag^+
- ions in a solution of silver acetate.

$$1) \text{Ag}^+ \text{ is } 1.10 \times 10^{-2} \text{ moles/liter} \quad \text{Ag}^+ \text{ is } 1.10 \times 10^{-2} \text{ moles/liter} \quad \text{Ag}^+ \text{ is } 1.10 \times 10^{-2} \text{ moles/liter}$$

2. How many moles of
- Ca^{2+}
- ions are there in 2.00 kg of
- CaCl_2
- ?

$$2) \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg} \quad \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg} \quad \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg}$$

3. What is the mass of 0.02 moles of
- Ca^{2+}
- ions? (molar mass of
- Ca^{2+}
- is 40.08 g/mol)

$$3) \text{Ca}^{2+} \text{ is } 0.02 \text{ moles} \quad \text{Ca}^{2+} \text{ is } 0.02 \text{ moles} \quad \text{Ca}^{2+} \text{ is } 0.02 \text{ moles}$$

4. Find the number of moles of ions in 2.00 kg of
- CaCl_2
- (molar mass of
- CaCl_2
- is 110.98 g/mol)

$$4) \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg} \quad \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg} \quad \text{CaCl}_2 \text{ is } \frac{2.00 \text{ kg}}{1000} = 2.00 \text{ kg}$$

5. If you start off with 0.02 moles of
- Ca^{2+}
- ions, what is the molar mass of this ion? Please show all your work.

$$5) \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{0.02 \text{ moles}} = 1.00 \text{ moles/liter} \quad \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{0.02 \text{ moles}} = 1.00 \text{ moles/liter}$$

6. What volume of 0.02 M
- Ca^{2+}
- solution contains 0.02 moles of
- Ca^{2+}
- ions?

$$6) \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{0.02 \text{ M}} = 1.00 \text{ liter} \quad \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{0.02 \text{ M}} = 1.00 \text{ liter}$$

7. What is the mass of 0.02 moles of
- Ca^{2+}
- ions? (molar mass of
- Ca^{2+}
- is 40.08 g/mol)

$$7) \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{1.00 \text{ liter}} = 0.02 \text{ moles/liter} \quad \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{1.00 \text{ liter}} = 0.02 \text{ moles/liter}$$

8. Calculate the molar mass of a gas of
- Ca^{2+}
- ions (molar mass of
- Ca^{2+}
- is 40.08 g/mol)

$$8) \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{1.00 \text{ liter}} = 0.02 \text{ moles/liter} \quad \text{Ca}^{2+} \text{ is } \frac{0.02 \text{ moles}}{1.00 \text{ liter}} = 0.02 \text{ moles/liter}$$

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Puja Mehta

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