

**CHAPTER 22 REVIEW***Organic Chemistry***MIXED REVIEW****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ A saturated organic compound
- (a) contains all single bonds.  
 (b) contains at least one double or triple bond.  
 (c) contains only carbon and hydrogen atoms.  
 (d) is quite soluble in water.
2. Arrange the following in order of increasing boiling point:
- \_\_\_\_\_ a. ethane  
 \_\_\_\_\_ b. pentane  
 \_\_\_\_\_ c. heptadecane
3. Recall that isomers in organic chemistry have identical molecular formulas but different structures and IUPAC names.
- \_\_\_\_\_ a. Two isomers must have the same molar mass. True or False?  
 \_\_\_\_\_ b. Two isomers must have the same boiling point. True or False?
4. Explain why hydrocarbons with only single bonds cannot form geometric isomers.

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

5. Write the IUPAC name for the following structural formulas:



# Mastering Chemistry Answers Ch 22

**Marti Anne Maguire, Paula Stiles**



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web electron configuration atomic radius pm aluminum 13  $1s^2 2s^2 2p^6 3s^2 3p^1$  118 silicon 14  $1s^2 2s^2 2p^6 3s^2 3p^2$  111 phosphorus 15  $1s^2 2s^2 2p^6 3s^2 3p^3$  98 sulfur 16  $1s^2 2s^2 2p^6 3s^2 3p^4$  88 chlorine 17  $1s^2 2s^2 2p^6 3s^2 3p^5$  79 argon 18  $1s^2 2s^2 2p^6 3s^2 3p^6$  71

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web photoelectron spectroscopy electron configurations google classroom you might need periodic table using  $s p d f$  notation what is the electron configuration for a neutral atom of beryllium choose 1 answer  $1s^2 2s^2$  a  $1s^2 2s^2 1s^2 2s^2 1s^2 2p^1$  b  $1s^2 2s^2 1s^2 2p^1$  c  $1s^2 2p^2$  d  $1s^2 2p^2$  e  $1s^2 2p^2 1s^2 2p^2$  show periodic table stuck

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an atomic number of 8 means a total of eight electrons two of these electrons can fit into the first shell while the remaining six can fit into the second shell therefore the configuration is written as 2 6

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web the four different types of orbitals s p d and f have different shapes and one orbital can hold a maximum of two electrons the p d and f orbitals have different sublevels thus can hold more electrons as stated the electron configuration of each element is unique to its position on the periodic table

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web explanation the first two electrons fit into the 1s orbital 1s<sup>2</sup> then next two into the 2s orbital 2s<sup>2</sup> the 2p orbital has three slots which can take 2 electrons each a total of 6 which gives 2p<sup>6</sup> the final two electrons fit in the 3s orbital 3s<sup>2</sup> electron configuration guided notes teacher edition

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web using the periodic table to determine the electron configurations of atoms is key but also keep in mind that there are certain rules to follow when assigning electrons to different orbitals the periodic table is an incredibly helpful tool in writing electron configurations

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$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^6$  ruthenium 23 kr  $5s^2 4d^{10}$  cadmium 24 xe  $6s^2 4f^{14} 5d^{10} 6p^2$  lead 25 rn  $7s^2 5f^{14} 6d^4$

seaborgium determine if the following electron configurations are correct 26  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^6 5s^1$  no it should be  $3d^{10}$

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web how many protons neutrons and electrons are in atoms of these isotopes write the complete electron configuration for each isotope answer co has 27 protons 27 electrons and 33 neutrons  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$  i has 53 protons 53 electrons and 78 neutrons  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^5$

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